

1. How do the bond angles in CH_4 , NH_3 and H_2O compare?

- A. $\text{CH}_4 = \text{NH}_3 = \text{H}_2\text{O}$
- B. $\text{CH}_4 < \text{NH}_3 < \text{H}_2\text{O}$
- C. $\text{NH}_3 < \text{CH}_4 < \text{H}_2\text{O}$
- D. $\text{H}_2\text{O} < \text{NH}_3 < \text{CH}_4$

2.

The electronegativities of four different elements are given below (the letters are not their chemical symbols).

Element	W	X	Y	Z
Electronegativity	0.9	1.2	3.4	4.0

Based on this information which statement is correct?

- A. W is a non-metal.
- B. W and X form an ionic compound.
- C. Y is a metal.
- D. Y and Z form a covalent compound.

3.

Which properties of the alkali metals decrease going down group 1?

- A. First ionization energy and reactivity
- B. Melting point and atomic radius
- C. Reactivity and electronegativity
- D. First ionization energy and melting point

2013 Ib Chemistry Exam Answers

Mary Louise Turgeon



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