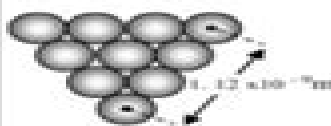

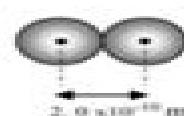
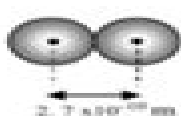


WORKSHEET FOUR

RADIi OF ATOMS AND IONS

1. (a) Determine the atomic radius of the four elements shown below by using the information given in the chart.

Element (not to scale)	Metallic sodium	Phosphorus (P ₄)	Chlorine (Cl ₂)	Iodine (I ₂)
				
Atomic radius				

- (b) Sodium, silicon and chlorine are all in the same row of the Periodic table.

(i) Describe the trend in atomic radii across a period.

(ii) Explain the reasons for this trend.

- (c) Chlorine and iodine are in the same group of the Periodic table.

(i) Describe the trend in atomic radii down a group.

(ii) Explain the reasons for this trend.

- (d) Explain why it is difficult to measure the atomic radii of the "Inert Gas" elements.

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George Graybill



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