

Question	Answer
1	increase in temperature 15 (°C)
2	energy released 20 = 4.2 × 15 = 630 (J)
3	energy released by 1 g 1260 = 6.30 × 4000 (J)

Question	Answer
4	An answer that makes reference to five of the following points: • student value much lower / eq (1) • not all energy released from food (1) • not burnt in oxygen / not completely combustion (1) • energy / heat not all transferred to tube / some energy / heat lost (1) to atmosphere (1) eq (1) • energy lost / heat to atmosphere when moved food / flame (1) • energy lost as light / eq (1)

The table shows some of the student's results.

Initial temperature of water in °C	Final temperature of water in °C	Increase in water temperature in °C	Energy released in J	Mass of bread sample in g	Energy released by 1 g of bread in J
20	35	15	1890	0.25	4000
20	34	14	1176	0.30	3920
21	36	15	1890	0.25	4000

(a) To calculate the energy released from the burning bread the student used this formula:

$$\text{energy (in J)} = \text{mass of water (in g)} \times 4.2 \times \text{increase in temperature (in } ^\circ\text{C)}$$

They then calculate the energy in joules released by 1 g of bread.

(i) Calculate the increase in water temperature for the third sample of bread.

temperature increase = _____ °C

(ii) Calculate the energy released in joules by the third sample of bread.

1000 g of water has a mass of 1 g = 1000

$$15 \times 4.2 \times 20$$

energy released = _____ J

(iii) Calculate the energy released in joules by 1 g of bread for the third sample of bread.

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Jicheng Xie



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