

**25**

# NUCLEAR CHEMISTRY

## Practice Problems

In your notebook, solve the following problems.

### SECTION 25.1 NUCLEAR RADIATION

- What happens to the mass number and atomic number of an atom that undergoes beta decay?
- A nuclide of an element undergoes alpha particle decay. How do the atomic number and mass number of the particle change?
- Give the composition of the nucleus of the following isotopes.
  - $^{224}_{86}\text{Rn}$
  - $^{232}_{90}\text{Th}$
  - $^{209}_{83}\text{Bi}$
- Complete each of the following equations.
  - $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th} + ?$
  - $^{226}_{88}\text{Ra} \rightarrow ^{222}_{86}\text{Rn} + ?$
  - $^{234}_{90}\text{Th} \rightarrow ^{234}_{91}\text{Pa} + ?$

### SECTION 25.2 NUCLEAR TRANSFORMATIONS

- Write a nuclear equation for the following radioactive processes.
  - alpha decay of francium-223
  - electron capture by beryllium-7
  - beta emission by argon-37
  - positron emission by boron-12
- Complete the equations for these transmutation reactions.
  - $^{24}_{12}\text{Mg} + ^4_2\text{He} \rightarrow ^{28}_{14}\text{Si} + ?$
  - $^{238}_{92}\text{U} + ^4_2\text{He} \rightarrow ? + ^{232}_{90}\text{Th} + ^1_0\text{n}$
  - $^{240}_{94}\text{Pu} + ^1_0\text{n} \rightarrow ? + ^{140}_{54}\text{Xe}$
  - $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th} + ? + ^4_2\text{He}$
  - $^9_4\text{Be} + ? \rightarrow ^{12}_6\text{C} + ^1_0\text{n}$
- Po-214 has a relatively short half-life of 164 s. How many micrograms would it take for 0.05 g of this isotope to decay to 0.25 g?
- How many days does it take for 10. g of gadolinium-153 to decay to 1.0 g? The half-life of gadolinium-153 is 17 days.
- By approximately what factor would the mass of a sample of copper-66 decrease in 50 minutes? The half-life of copper-66 is 5.48 min.
- In 5.48 minutes, 1.00 g of argon-39 decays to leave only 0.10 g. What is the half-life of argon-39?

## 5 Nuclear Radiation Answer Key

**Prentice-Hall Staff**

A red circular graphic with a gradient, appearing as a partial circle or a stylized 'C' shape, located to the right of the Prentice-Hall Staff text.

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