

# Practice 4-1

Complementary Angles

Two angles are **complementary** if the sum of their measures is 90°. Two angles are **supplementary** if the sum of their measures is 180°. In each pair of polygons below, identify the pair of angles that are complementary or supplementary. Label the measures of the unlabeled angles.



**Set A** or **Set B**, whichever of the following:

1. three pairs of complementary angles
2. three pairs of supplementary angles



**Set C** or **Set D**, whichever of the following:

1. three pairs of complementary angles
2. three pairs of supplementary angles



Now identify the pairs of angles are complementary or supplementary.

1.  $\angle ABC$  and  $\angle DEF$

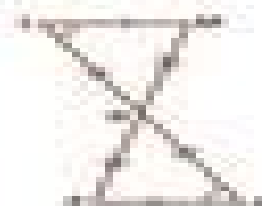


2.  $\angle GHI$  and  $\angle JKL$



3. Identifying Pairs of Angles: Given the diagram, identify the pairs of angles that are complementary or supplementary.

- a.  $\angle A$  and  $\angle B$
- b.  $\angle C$  and  $\angle D$
- c.  $\angle E$  and  $\angle F$
- d.  $\angle G$  and  $\angle H$ ,  $\angle I$  and  $\angle J$ ,  $\angle K$  and  $\angle L$
- e.  $\angle M$  and  $\angle N$



# 4 1 Practice Form G Congruent Figures

**Wolfgang Guggemos**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients  $\text{Cu} + \text{MnO}_4^- + 4\text{H}^+ \rightarrow \text{Cu}^{2+} + \text{MnO}_2 + 2\text{H}_2\text{O}$  is

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the following reduction oxidation reaction in basic solution what is the coefficient in front of oh br 2 l bro 3 aq br aq previous  
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web google classroom for the following redox equation  $\text{MnO}_4^- + \text{H}_2\text{SO}_3 \rightarrow \text{Mn}^{2+} + \text{HSO}_4^-$  what will be the stoichiometric coefficient of  $\text{HSO}_4^-$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow 2\text{e}^- + 2\text{H}^+ + \text{O}_2 + \text{H}_2\text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{OH}^- \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 + \text{H}_2\text{O}_2 + 2\text{OH}^-$  the final answer

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basic solution created by jay

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution  $4\text{Fe}^{2+} + 5\text{H}_2\text{O} + 6\text{OH}^- \rightarrow 4\text{Fe}^{3+} + 2\text{H}_2\text{O} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method  $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te} + 10\text{H}_2\text{O}$   $11\text{PbO}_2 + \text{IO}_2 + 2\text{Pb} \rightarrow 12\text{As} + \text{NO}_3^- + \text{IO}_3^- + \text{H}_2\text{AsO}_4$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  no 2 and  $\text{C}_6\text{H}_8\text{O}_6$  b  $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  h and no 2 c h and