

Chemistry 20

Unit C: Matters as Solutions, Acids and Bases “I Can” Statements

CB1	Students will describe theoretically and qualitatively solution formation and properties of solutions.
a.	I can recall the sequence of steps, solutes and solvents and explain the nature of homogeneous mixtures.
b.	I can provide examples from living and nonliving systems that illustrate how dissolving substances in water is often a prerequisite for chemical change.
c.	I can explain dissolving as an endothermic or exothermic process with respect to the breaking and forming of bonds.
d.	I can differentiate between electrolytes and nonelectrolytes.
CB2	Students will express and calculate concentrations using molarity, percent, and ppm.
a.	I can express concentration in various ways, i.e., mole per litre of solution, percent by mass and parts per million.
CB3	Students will perform calculations involving the preparation of solutions from a solid and in diluting.
a.	I can calculate, from empirical data, the concentration of solutions in mole per litre of solution and determine mass or volume from such concentrations.
b.	I can calculate the concentrations and/or volumes of diluted solutions and the quantity of a solution and water to use when diluting.
c.	I can describe the procedures and calculations required for preparing and diluting solutions.
CB4	Students will quantitatively and qualitatively represent solution formation including dissociation equations, ionization equations, and acid ionization and ion concentrations.
a.	I can use data and ionization/dissociation equations to calculate the concentration of ions in a solution.
b.	I can define solubility and identify related factors, i.e., temperature, pressure and ionicity.
c.	I can explain a saturated solution in terms of equilibrium, i.e., equal rates of dissolving and crystallization.
CB5	Students will accurately describe substances using formulas, structural formulas, and IUPAC naming.
a.	I can recall the structural formula of HCl and Applied Chemistry (ACAC) nomenclature of acids and bases.
b.	I can recall the empirical definitions of acids, bases and neutral solutions determined by using indicators, pH and electrical conductivity.
CB6	Students will describe acid and base forming in weak, monoprotic or polyprotic using the modified Arrhenius theory and explain other water sharing neutralization reactions.
a.	I can define Arrhenius (modified) acids as substances that produce H_3O^+ in aqueous solutions and recognize that this definition is limited.
b.	I can define Arrhenius (modified) bases as substances that produce OH^- in aqueous solutions and recognize that this definition is limited.
c.	I can define neutralization as a reaction between hydronium and hydroxide ions.
d.	I can differentiate, qualitatively, between strong and weak acids and between strong and weak bases on the basis of ionization and dissociation, i.e., pH, reaction rate and electrical conductivity.
e.	I can identify monoprotic and polyprotic acids and bases and compare their ionization/dissociation.
CB7	Students will calculate pH, pOH, $[H_3O^+]$ and $[OH^-]$ and compare magnitude changes when comparing solutions with different concentrations.
a.	I can calculate $[H_3O^+]$ and $[OH^-]$ concentrations and the pH and pOH of acidic and basic solutions based on logarithmic equations, i.e., $pH = -\log[H_3O^+]$ and $pOH = -\log[OH^-]$.
b.	I can use appropriate SI units to communicate the concentration of solutions and express pH and concentration answers to the correct number of significant digits, i.e., use the number of decimal places in the pH to determine the number of significant digits of the concentration.
c.	I can compare magnitude changes in pH and pOH with changes in concentration for acids and bases.
CB8	Students will describe how acid/base/indicator can be measured using indicators.
a.	I can explain how the use of indicators, pH and pOH with changes in concentration for acids and bases.

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