

Chapter 7 Chemical Reactions

Section 7.1 Describing Reactions

(pages 192–198)

This section discusses the use of chemical equations and how to balance them. It also demonstrates the use of calculations in chemistry.

Reading Strategy (page 192)

Monitoring Your Understanding Preview the Key Concepts, topic headings, vocabulary, and figures in this section. List two things you expect to learn. After reading, state what you learned about each item you listed. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Expect to Learn	What I Learned
Answers may vary. Possible answers: How to balance chemical equations	Answers may vary. Possible answers: An unbalanced equation can be balanced by changing the coefficients.
How to convert from mass to moles	The mass of a substance can be converted to moles by using the molar mass as a conversion factor.

Chemical Equations (pages 192–193)

- Is the following sentence true or false? The new substances formed as a result of a chemical reaction are called products. true
- Circle the letter of each sentence that is a correct interpretation of the chemical equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$.
 - Carbon and oxygen react and form carbon monoxide.
 - ☒ Carbon and oxygen react and form carbon dioxide.
 - Carbon dioxide yields carbon and oxygen.
 - ☒ The reaction of carbon and oxygen yields carbon dioxide.
- Is the following sentence true or false? The law of conservation of mass states that mass is neither created nor destroyed in a chemical reaction. true
- Circle the letter of the correct answer. According to the equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$, how many carbon atoms react with 14 molecules of oxygen to form 14 molecules of carbon dioxide?
 - 1
 - 7
 - ☒ 14
 - 28
- In the reaction represented by the equation $\text{C} + \text{O}_2 \longrightarrow \text{CO}_2$, the mass of carbon dioxide produced equals the total mass of carbon and oxygen that reacted.

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Table of Contents 7 1 Section Review Pearson Education Answer Key

1. Understanding the eBook 7 1 Section Review Pearson Education Answer Key
 - The Rise of Digital Reading 7 1 Section Review Pearson Education Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 7 1 Section Review Pearson Education Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 7 1 Section Review Pearson Education Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 7 1 Section Review Pearson Education Answer Key
 - Personalized Recommendations
 - 7 1 Section Review Pearson Education Answer Key User Reviews and Ratings
 - 7 1 Section Review Pearson Education Answer Key and Bestseller Lists

5. Accessing 7 1 Section Review Pearson Education Answer Key Free and Paid eBooks
 - 7 1 Section Review Pearson Education Answer Key Public Domain eBooks
 - 7 1 Section Review Pearson Education Answer Key eBook Subscription Services
 - 7 1 Section Review Pearson Education Answer Key Budget-Friendly Options
6. Navigating 7 1 Section Review Pearson Education Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 7 1 Section Review Pearson Education Answer Key Compatibility with Devices
 - 7 1 Section Review Pearson Education Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 7 1 Section Review Pearson Education Answer Key
 - Highlighting and Note-Taking 7 1 Section Review Pearson Education Answer Key
 - Interactive Elements 7 1 Section Review Pearson Education Answer Key
8. Staying Engaged with 7 1 Section Review Pearson Education Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 7 1 Section Review Pearson Education Answer Key
9. Balancing eBooks and Physical Books 7 1 Section Review Pearson Education Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 7 1 Section Review Pearson Education Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 7 1 Section Review Pearson Education Answer Key
 - Setting Reading Goals 7 1 Section Review Pearson Education Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 7 1 Section Review Pearson Education Answer Key
 - Fact-Checking eBook Content of 7 1 Section Review Pearson Education Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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