

7-2 Review and Reinforcement

Covalent Bonding

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line provided.

_____ F, covalent 1. A group of atoms united by ionic bonds is called a molecule.

_____ T 2. A covalent bond is formed by a shared pair of electrons

_____ F, four 3. A double covalent bond consists of two shared electrons.

_____ T 4. A molecular formula tells you how many atoms are in a single molecule of the compound.

_____ F, MF 5. The empirical formula for a molecule specifies which atoms are bonded to each other in the molecule.

_____ F, lone pair 6. A pair of electrons not involved in bonding is called a shared electron pair.

_____ F, equally 7. In nonpolar covalent bonds, the electrons are shared unequally between two atoms.

_____ F, Triple 8. Lewis structures use a triple dash to represent a double bond.

Use the information in Figure 7-19 of the textbook to calculate the electronegativity differences between each pair of elements listed in the chart below. Then indicate whether the bond would be nonpolar covalent, polar covalent, or ionic. Write your answers in the chart.

		Electronegativity Differences	Type of Bond
9.	H-O		
10.	C-H		
11.	K-F		
12.	N-H		
13.	Na-F		
14.	O-Cl		

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Table of Contents 7 2 Review And Reinforcement

1. Understanding the eBook 7 2 Review And Reinforcement
 - The Rise of Digital Reading 7 2 Review And Reinforcement
 - Advantages of eBooks Over Traditional Books
2. Identifying 7 2 Review And Reinforcement
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 7 2 Review And Reinforcement
 - User-Friendly Interface
4. Exploring eBook Recommendations from 7 2 Review And Reinforcement
 - Personalized Recommendations
 - 7 2 Review And Reinforcement User Reviews and Ratings
 - 7 2 Review And Reinforcement and Bestseller Lists
5. Accessing 7 2 Review And Reinforcement Free and Paid eBooks
 - 7 2 Review And Reinforcement Public Domain eBooks
 - 7 2 Review And Reinforcement eBook Subscription Services
 - 7 2 Review And Reinforcement Budget-Friendly Options

6. Navigating 7 2 Review And Reinforcement eBook Formats
 - ePub, PDF, MOBI, and More
 - 7 2 Review And Reinforcement Compatibility with Devices
 - 7 2 Review And Reinforcement Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 7 2 Review And Reinforcement
 - Highlighting and Note-Taking 7 2 Review And Reinforcement
 - Interactive Elements 7 2 Review And Reinforcement
8. Staying Engaged with 7 2 Review And Reinforcement
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 7 2 Review And Reinforcement
9. Balancing eBooks and Physical Books 7 2 Review And Reinforcement
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 7 2 Review And Reinforcement
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 7 2 Review And Reinforcement
 - Setting Reading Goals 7 2 Review And Reinforcement
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 7 2 Review And Reinforcement
 - Fact-Checking eBook Content of 7 2 Review And Reinforcement
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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