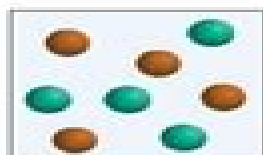


5	Caryl Chan	11-Service	October 17, 2020	
Class #	Name	Section	Date	Score

**Precipitation reaction**

**Directions:** The aqueous solutions of 3 compounds are shown below. Identify which is a non-electrolyte, a weak electrolyte and a strong electrolyte. Explain your answer



(a)



(b)



(c)

a. Strong electrolyteb. Non-electrolytec. Weak electrolyte**Explanation:**

The first solution is the strongest electrolyte because it completely ionizes. And when it completely ionizes, the molecules are more spread out which is evident in the placement of the molecules. The last solution is the weak electrolyte because the molecules do not completely ionize. This is evident in the way some molecules are bonded and some are not. And the solution in the middle is a non-electrolyte because the molecules do not move at all.

**Directions:** Identify if the following compounds are soluble or insoluble in water according to the solubility rules. Draw a (+) if the compound is soluble and draw an (x) if the substance is insoluble.

1.  $\text{Ca}_3\text{PO}_4$  x4.  $\text{K}_2\text{S}$  +6.  $\text{Hg}_2\text{NO}_3$  +2.  $\text{BaCO}_3$  x5.  $\text{ZnSO}_4$  +7.  $\text{Hg}_2\text{Cl}_2$  x3.  $\text{AgCl}$  x8.  $\text{NH}_4\text{OH}$  +

## 2 4 Precipitation Reactions Answers

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