

Chemistry: Chapter 1 & 2 Review

Name: _____

1. A mixture is is not a chemical combination of substances.
2. In a compound the atoms (molecules) are chemically physically combined so that the elements that make up the compound retain the same identities and do (do not) take on new set of properties.
3. The smallest identifiable unit of a compound is also MOLECULE, which is made up of TWO OR MORE ATOMS, which are chemically bonded.
4. True or False A mixture is always made up of a combination of elements.
COULD BE A COMBINATION OF COMPOUNDS
5. In a mixture, the substances (have retain) their identities.
6. In a mixture, the substances involved can cannot be separated by a simple physical process.
7. In a compound, the elements involved (are cannot) be separated by a simple physical process because the elements are physically combined chemically bonded.
8. The smallest identifiable unit of an element is a(n) ATOM.
9. Explain how to separate the sugar and water in a solution of sugar and water.
BOIL (OR EVAPORATE) THE WATER OFF, LEAVING THE SUGAR BEHIND.
10. How would you separate a mixture of alcohol and water?
DISTILLATION -- HEAT MIXTURE TO THE 1st BOILING POINT (ALCOHOL)
Collect Vapors --> Condenser --> Liquid
11. How would you separate sand and water?
FILTER -- WATER MIXTURE PASSES THRU LEAVING THE SAND IN THE FILTER
12. Classify the following as pure substances or as mixtures:

Gasoline	<u>S.M</u>	Sugar	<u>S</u>	Air	<u>M</u>
Mercury	<u>S</u>	Salt water	<u>M</u>	Oxygen	<u>S</u>
Distilled Water	<u>S</u>	Blood	<u>M</u>	Nitrogen Gasoline	<u>M</u>

13. What is the difference between Extensive and Intensive Properties?
14. Why are "intensive properties" so important?

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Morris Hein,Susan Arena

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