

## Factors affecting electronic conductance

- Nature of conductor
- No. of valence of e
- Temp ↑ K ↓

## Factors affecting electrolytic or ionic conductance

① Nature  
Strong electrolyte → more

② Nature of solvent  
high polarity  
high will be dissociation

③ Temp  
T ↑ KE ↑ mobility ↑ K ↑

## FOR NUMERICAL'S

$$R = \frac{\rho \cdot l}{A} \quad \Omega$$

$$G^* = \frac{\sigma}{l} \quad G^* = \frac{K}{m} \quad m^{-1}$$

$$\rho = \frac{1}{K} \quad \frac{\Omega \cdot m}{S \cdot m^{-1}}$$

$$G = \frac{1}{R} \quad \frac{S}{moh \cdot simen}$$

## molar conductivity

$$\lambda_m = \frac{K \times 1000}{M} \quad S \cdot cm^2 \cdot mol^{-1}$$

$$\lambda_m = \frac{K}{1000 M} \quad S \cdot m^2 \cdot mol^{-1}$$

## Equivalent conductivity

$$\lambda_{eq} = \frac{K \times 1000}{N} \quad S \cdot cm^2 \cdot eq^{-1}$$

$$\lambda_{eq} = \frac{K}{N \times 1000} \quad S \cdot m^2 \cdot eq^{-1}$$

Ch=2 Electrochemistry

## Variation with conc

K always ↓ conc ↓ V ↑  
(we, se) (volume)

λ<sub>m</sub> ↑ conc ↓ V ↑  
as conc ↓  
interionic attraction ↓

Strong electrolyte  
λ<sub>m</sub> increase slowly with dilution  
λ<sub>m</sub> = λ<sub>m</sub><sup>∞</sup> - A C<sup>1/2</sup>

Weak electrolytes  
for WE lower α at higher conc  
on dilution α ↑ V ↑

for WE, λ<sub>m</sub><sup>∞</sup> measured by Kohlrausch



## effect of dilution

dilution ↑ Volume ↑  
dilution ↑ K ↓ Volume ↑  
dilution ↑ λ<sub>m</sub> ↑ Volume ↑

→ magnitude of inc. in molar conductance for weak electrolyte is much larger than that for SE becoz Strong electrolyte completely ionised in all conc

## Kohlrausch law

independent migration of ions

$$\alpha = \frac{\lambda_m}{\lambda_m^{\infty}}$$

$$K_0 = \frac{C \alpha^2}{(1 - \alpha)}$$

law state that limiting molar conductivity of an electrolyte can be represented as the sum of the individual contribution of anion of the electrolyte

$$\lambda_m^{\infty} = \nu_+ \lambda_c + \nu_- \lambda_a$$

## Electrolysis

1st law

$$W = ZQ$$

$$W = \frac{E}{F} Q$$

$$W = \frac{m}{nF} Q$$

$$W = \frac{m}{nF} It$$

$$F = 96500$$

$$\frac{W}{m} \times nF = \frac{It}{F} = \frac{Q}{F}$$

$$n \times nF = \frac{It}{F} = \frac{Q}{F}$$

$$V \times nF = n \times nF = \frac{It}{F}$$

$$22.4$$

## Note

equivalence = N X V

$$= \frac{W}{F}$$

$$= n \times nF$$

$$= \frac{It}{F}$$

## Faraday II Law

equal charge passed through diff electrolytes amount of substance deposited is proportional to equivalent wt

$$\frac{W_1}{W_2} = \frac{Z_1}{Z_2} = \frac{E_1}{E_2}$$

## Electrochemical series

SRP ↑ Reduction cathode oxidising agent

SRP ↓ oxidation anode Reducing agent

## Reactivity of metals

Reactivity α 1/SRP

P > C > S > M > A > Z > N > D > H  
K > Ca > Na > Mg > Al > Zn > Fe > Sn > Pb

C > H > A > R > P  
Cu > Hg > Ag > Au > Pt

## Reactivity of non metals

Reactivity α SRP

F > S > N > O > Cl > Br > I

## Potential discharge theory

E < S < M < A < Z < N < D < H

## EMF

E < S < N < D < Cl < Br < I

ions having less discharge potential will release first

## IMP FORMULA

$$E^{\circ} = E_c^{\text{SRP}} - E_a^{\text{SRP}}$$

$$E = E^{\circ} - \frac{0.059}{n} \log \frac{P}{R}$$

$$\Delta G^{\circ} = -nE^{\circ}F$$

$$E^{\circ} = \frac{0.06}{n} \log K$$

$$\Delta G^{\circ} = -2.303 nRT \log K$$

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**Sir Norman Lockyer**



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