

	In the following questions one mark each (Q. No. 6 - 9) a statement of Assertion (A) followed by a statement of Reason (R) is given. Choose the correct answer out of the following choices. (i) A and R both are correct statements and R is the correct explanation for A. (ii) A and R both are correct statements and R is not correct explanation for A. (iii) A is correct statement but R is wrong statement. (iv) A is wrong statement but R is correct statement.	
6.	Assertion (A): The empirical mass of ethene is half of its molecular mass. Reason (R): The empirical formula represents the simplest whole-number ratio of various atoms present in a compound.	1
7.	Assertion (A): Equal moles of different substances contain same number of constituent particles. Reason (R): Equal weights of different substances contain the same number of constituent particles.	1
8.	Assertion: It is impossible to determine the exact position and exact momentum of an electron simultaneously. Reason: The path of an electron in an atom is clearly defined.	1
9.	Assertion: The radius of the first orbit of hydrogen atom is $0.529\text{\AA}$. Reason: Radius of each circular orbit (r_n) = $0.529\text{\AA} (n^2/Z)$, where $n = 1, 2, 3$ and Z = atomic number.	1
10.	45.4 L of dinitrogen reacted with 22.7 L of dioxygen and 45.4 L of nitrous oxide was formed. The reaction is given below: $2\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{N}_2\text{O}(\text{g})$ Which law is being obeyed in this experiment? Write the statement of the law.	2/4
11.	The density of the 3 molal solution of NaOH is 1.110 g mL^{-1} . Calculate the molarity of the solution.	2
12.	What is the difference between the terms orbit and orbital?	2
13.	A cricket ball of mass 100 g does not move like a wave when it is thrown by a bowler at a speed of 100 km/h. Calculate the wavelength of the ball and explain why it does not show wave nature.	2
14.	Calculate the mass percent of calcium, phosphorus and oxygen in calcium phosphate $\text{Ca}_3(\text{PO}_4)_2$	3
15.	(a) 45.4 L of dinitrogen reacted with 22.7 L of dioxygen and 45.4 L of nitrous oxide was formed. The reaction is given below: $2\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{N}_2\text{O}(\text{g})$. Which law is being obeyed in this experiment? Write the statement of the law. (b) How many significant figures should be present in answer to the following calculation? $\frac{2.5 \times 1.25 \times 3.5}{2.01}$	3

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