## Balancing Equations Worksheet II

Count the number of atoms, then use coefficients to balance each equation.

AlBr<sub>3</sub> + \_\_\_ K → \_\_\_ KBr + \_\_\_ Al

2) \_\_\_\_ FeO + \_\_\_ PdF₂ → \_\_\_ FeF₂ + \_\_\_ PdO

3) \_\_\_\_ P<sub>4</sub> + \_\_\_\_ Br<sub>2</sub> → \_\_\_\_ PBr<sub>3</sub>

4) \_\_\_\_ LiCl + \_\_\_ Br<sub>2</sub> → \_\_\_ LiBr + \_\_\_ Cl<sub>2</sub>

5) \_\_\_\_ PbBr<sub>2</sub> + \_\_\_ HCl → \_\_\_ HBr + \_\_\_ PbCl<sub>2</sub>

6) \_\_\_\_ CoBr<sub>3</sub> + \_\_\_\_ CaSO<sub>4</sub> → \_\_\_\_ CaBr<sub>2</sub> + \_\_\_\_ Co<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>

7) \_\_\_\_ Na<sub>3</sub>P + \_\_\_ CaF<sub>2</sub> → \_\_\_ NaF + \_\_\_ Ca<sub>3</sub>P<sub>2</sub>

8) \_\_\_\_ Mn + \_\_\_ HI → \_\_\_ H₂ + \_\_\_ MnI₃

9) \_\_\_\_ Li<sub>3</sub>PO<sub>4</sub> + \_\_\_\_ NaBr → \_\_\_\_ Na<sub>3</sub>PO<sub>4</sub> + \_\_\_\_ LiBr

10) \_\_\_\_ CaF<sub>2</sub> + \_\_\_\_ Li<sub>2</sub>SO<sub>4</sub> → \_\_\_\_ CaSO<sub>4</sub> + \_\_\_\_ LiF

11) \_\_\_\_ HBr + \_\_\_ Mg(OH)<sub>2</sub> → \_\_\_ MgBr<sub>2</sub> + \_\_\_ H<sub>2</sub>O

12) \_\_\_\_ LiNO<sub>3</sub> + \_\_\_\_ CaBr<sub>2</sub> → \_\_\_\_ Ca(NO<sub>3</sub>)<sub>2</sub> + \_\_\_\_ LiBr

13) \_\_\_\_ AgNO<sub>3</sub> + \_\_\_ Li → \_\_\_ LiNO<sub>3</sub> + \_\_\_ Ag

14) \_\_\_\_ Si(OH)₄ + \_\_\_ NaBr → \_\_\_ SiBr₄ + \_\_\_ NaOH

15) \_\_\_\_ NaCN + \_\_\_ CuCO<sub>3</sub> → \_\_\_ Na<sub>2</sub>CO<sub>3</sub> + \_\_\_ Cu(CN)<sub>2</sub>

Challenge: \_\_\_\_  $C_2H_6 + ___ O_2 \rightarrow ___ CO_2 + ___ H_2O$ 

# **1chemical Reaction Practice Problems Answer Key**

**Morris Hein, Susan Arena** 

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## **Table of Contents 1chemical Reaction Practice Problems Answer Key**

- 1. Understanding the eBook 1chemical Reaction Practice Problems Answer Key
  - The Rise of Digital Reading 1chemical Reaction Practice Problems Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying 1chemical Reaction Practice Problems Answer Key
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 1chemical Reaction Practice Problems Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from 1chemical Reaction Practice Problems Answer Key
  - Personalized Recommendations
  - 1chemical Reaction Practice Problems Answer Key User Reviews and Ratings
  - 1chemical Reaction Practice Problems Answer Key and Bestseller Lists

- 5. Accessing 1chemical Reaction Practice Problems Answer Key Free and Paid eBooks
  - 1chemical Reaction Practice Problems Answer Key Public Domain eBooks
  - 1chemical Reaction Practice Problems Answer Key eBook Subscription Services
  - 1chemical Reaction Practice Problems Answer Key Budget-Friendly Options
- 6. Navigating 1chemical Reaction Practice Problems Answer Key eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - 1chemical Reaction Practice Problems Answer Key Compatibility with Devices
  - 1chemical Reaction Practice Problems Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 1chemical Reaction Practice Problems Answer Key
  - Highlighting and Note-Taking 1chemical Reaction Practice Problems Answer Key
  - Interactive Elements 1chemical Reaction Practice Problems Answer Key
- 8. Staying Engaged with 1chemical Reaction Practice Problems Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 1chemical Reaction Practice Problems Answer Key
- 9. Balancing eBooks and Physical Books 1chemical Reaction Practice Problems Answer Key
  - $\circ$  Benefits of a Digital Library
  - Creating a Diverse Reading Collection 1chemical Reaction Practice Problems Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 1chemical Reaction Practice Problems Answer Key
  - Setting Reading Goals 1chemical Reaction Practice Problems Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 1chemical Reaction Practice Problems Answer Key
  - Fact-Checking eBook Content of 1chemical Reaction Practice Problems Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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