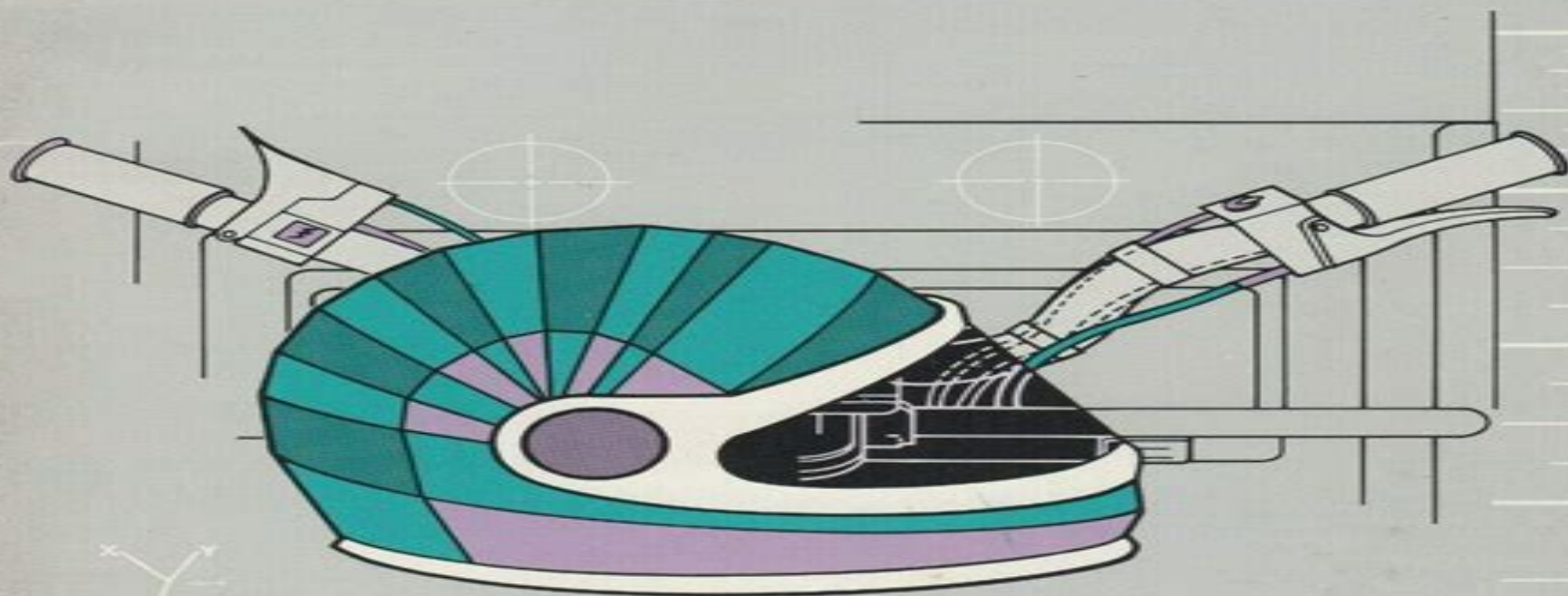


# operator's manual

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# 199ski Doo Mach Manual

**Annelies Wilder-Smith**



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web start with 8 3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8 3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149 4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137 1 \text{ g } \text{H}_2\text{O}}{149 4 \text{ g } \text{H}_2\text{O}} \times 100 = 91 77\%$  problems 6 8 are similar to problem 5

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web 7 Na 2 C 2 O 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of  $\text{H}_2$  are left over 13

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