

# U.S. ARMY ROTC 4-YEAR SCHOLARSHIP APPLICATION REQUEST FOR SECONDARY SCHOOL TRANSCRIPT

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## DATA REQUIRED BY THE PRIVACY ACT OF 1974

<b>Authority</b> <b>Principal Purpose</b> <b>Routine Uses</b>  <b>Disclosure</b>	10 USC 2101, 2103, 2104, 2107, 2107A, 2111 and 5 USC 301.2 To request applicant's secondary school transcript for use in determining scholarship eligibility. Capturing scholarship applicant's secondary school progress. Once completed by the school guidance counselor, this form is used in determining the applicant's eligibility for a scholarship. Disclosure of the information requested in this endorsement is voluntary. However, applicable portions must be completed if the applicant desires to be considered for an Army Senior ROTC Scholarship.
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## PERSONAL DATA

1. NAME OF APPLICANT (Last, First, Middle)	2. SOCIAL SECURITY NUMBER (Last 4 digits)	
3. BIRTH DATE (YY MM DD)	4. YEAR OF GRADUATION	5. HIGH SCHOOL NAME AND ETS CODE

The student named above is applying for an Army ROTC 4-Year Scholarship. Please complete this part of the form as accurately as possible. A transcript of grades is required in reviewing the applicant's record to determine scholarship eligibility.

A. In addition to courses taken (or in progress) and grades received, it is essential that the transcript reflect rank in class along with the most complete academic record to include test results such as ACT's, SAT's, and other national examinations.

B. **IMPORTANT!** Please submit this information immediately.

C. Return completed form and a transcript signed or stamped by a high school official to: Headquarters, Cadet Command ATTN: Incentives Division (Scholarships), 55 Patch Rd., Bldg. 56, Fort Monroe, VA 23651-1052.

D. I authorize the release of my high school transcript.

APPLICANT SIGNATURE \_\_\_\_\_

DATE \_\_\_\_\_

## SCHOLASTIC/BACKGROUND DATA

6. CUMULATIVE GRADE POINT AVERAGE <i>(To nearest tenth. Please convert to a 4.0 GPA Scale)</i>  8. % GRAD CLASS EXPECTED TO ENTER <div style="display: flex; justify-content: space-between;"> <div style="width: 40%;"> <input type="text"/> 4-YEAR COLLEGE  <input type="text"/> 2-YEAR COLLEGE </div> <div style="width: 60%;">           7. RANK IN CLASS <i>(Approximate to nearest tenth from top)</i>  <div style="display: flex; justify-content: space-around; margin-bottom: 10px;"> <input type="text"/> EXACTLY <input type="text"/> APPROXIMATELY </div> <div style="display: flex; justify-content: space-between;"> RANK FROM TOP <input style="width: 50px;" type="text"/> NUMBER OF STUDENTS <input style="width: 50px;" type="text"/> </div> </div> </div>
--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------

9. Did this student take any of the following:	Not Offered	Yes	No
Honors Courses	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
Accelerated Courses	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
Advanced Placement Courses	<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>

10. Are all honors accelerated and advanced placement courses given extra credit in computing:		
	Yes	No
Rank in Class	<input type="checkbox"/>	<input type="checkbox"/>
Grade Averages	<input type="checkbox"/>	<input type="checkbox"/>

11. Official School Name _____			
School Address _____			
City _____	State _____	Zip Code _____	
School Telephone Number <i>(include area code)</i> _____			
School E-Mail Address _____			

# Amry Form Application

**JS Bruner**



**Amry Form Application:**

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In a digitally-driven world where displays reign great and immediate transmission drowns out the subtleties of language, the profound strategies and mental nuances hidden within words frequently move unheard. However, nestled within the pages of **Amry Form Application** a fascinating literary treasure sporting with natural emotions, lies an extraordinary quest waiting to be undertaken. Written by a talented wordsmith, this enchanting opus encourages readers on an introspective journey, softly unraveling the veiled truths and profound affect resonating within the fabric of each and every word. Within the emotional depths of the touching evaluation, we can embark upon a genuine exploration of the book's core styles, dissect its captivating writing style, and succumb to the strong resonance it evokes serious within the recesses of readers' hearts.

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pride and prejudice and zombies is a novel by jane austen and seth graham smith and details the relationship between elizabeth bennet and mr darcy against the greater backdrop of a british zombie plague when the novel begins mrs bennet is determined to wed off her five daughters despite the zombie scourge mr

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 fe<sup>2+</sup> co co<sub>2</sub> fe oxidation co co<sub>2</sub> 2e reduction fe<sup>2+</sup> 2e fe 24 3 ag ni ni<sub>3</sub> 3 ag oxidation reduction 25 cu<sub>2</sub> pb pb<sub>2</sub> cu oxidation reduction 26 o 2

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 $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow \text{I}_2 + 5\text{CO}_2$

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 $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$  Sr 0 to Sr 2 oxidized reducing agent O 0 to O 2 reduced  
 $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$  Li 0 to Li 1 oxidized red ag S 0 to S 2 reduced  
 $3\text{C}_2\text{S} + 2\text{Br}_2 \rightarrow \dots$

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