

# Additivity of Heats of Reaction: Hess' Law

Goal = verify Hess's laws

In this experiment, you will use a Styrofoam-cup calorimeter to conserve energy so that the heat released by three reactions can be measured accurately. One of the reactions is the same as the combination of the other two reactions. Therefore, according to Hess's law, the heat of reaction of the one reaction should be equal to the sum of the heats of reaction for the other two. This concept is sometimes referred to as the *additivity of heats of reaction*. The primary objective of this experiment is to confirm this law. The reactions we will use in this experiment are:

- (1) Solid sodium hydroxide dissolves in water to form an aqueous solution of ions.



- (2) Solid sodium hydroxide reacts with aqueous hydrochloric acid to form water and an aqueous solution of sodium chloride.



- (3) Solutions of aqueous sodium hydroxide and hydrochloric acid react to form water and aqueous sodium chloride.



## OBJECTIVES

In this experiment, you will

- Combine equations for two reactions to obtain the equation for a third reaction.
- Use a calorimeter to allow for the temperature change in each of three reactions to be accurately measured.
- Calculate the heat of reaction,  $\Delta H$ , for the three reactions.
- Use the results to confirm Hess's law.

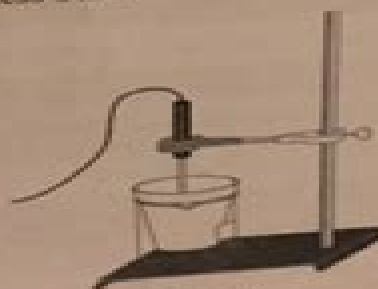


Figure 1

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**Karin Nielsen-Saines**



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