

EXPERIMENT 6: RELATIVE REACTIVITIES OF METALS AND THE ACTIVITY SERIES

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Instructor: Lee Jennings Section/Group: 322

Post-Lab Questions (Turn in to your instructor with the post laboratory report for this experiment.)

1. Based upon your observations, what general statement can you make about the reactivity of a metal and its relative position on the periodic table? (Where are the most reactive metals located in comparison to the less reactive metals?)

Reactivity decreases as you go from left to right on a periodic table.

2. It took many of the metals several minutes to begin to react. Once the reaction was initiated, several of these metals continued to react quite vigorously. What was the most probable cause of this delayed reaction? (Hint: The metal getting hot isn't the only reason.)

3. How would your ranking of the relative reactivity of metals have change if they were all first compared with hot aqueous HCl?

4. What are some of the errors that may have occurred in this experiment that may have caused you to assign a metal to the wrong position in your activity series?

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