

An Activity Series

Purpose: The purpose of this lab is to verify the activity series of five metals and three halogens. This can be done by reacting the metal with various metal solutions to see which of them creates a displacement reaction. As for the halogens, it can be done in a similar way as metal, but with the halogens in their liquid form and ion form. The significance of this experiment is that geologist could use them to extract minerals from ores.

Hypothesis: The hypothesis of this experiment is that the experiment will help verify the activity series of metals and halogens. This is because the reaction would not occur if the substance trying to replace the other substance is less reactive than that other substance. If the reaction did happen, there would be an apparent change in the sample. For example, if one metal was replaced by the other, the solid in the solution would be changed to the metal that was replaced. As for halogens, each halogen possess different color from another and so if one halogen was replaced by the other, the color of the solution would be changed.

Materials:

Part 1	
Materials	Quantity
6 × 6 mm Copper foil (Cu)	4 mm ³
0.1 M Copper (II) nitrate (Cu(NO ₃) ₂)	4mL
6 × 6 mm Lead foil (Pb)	4 mm ³
0.1 M Lead nitrate (Pb(NO ₃) ₂)	4mL
6 mm Magnesium ribbon (Mg)	4 mm ³
0.1 M Magnesium nitrate (Mg(NO ₃) ₂)	4mL
0.1 M Silver nitrate (AgNO ₃)	4mL
6 × 6 mm Zinc foil (Zn)	4 mm ³
0.1 M Zinc nitrate (ZnNO ₃)	4mL
5mL Beral-type pipette	5 pipettes
13 × 100mm test tube	24 test tubes
Labeling tape	30 labels

Part 2	
Materials	Quantity
Bromine water (Br ₂)	3mL
Chlorine water (Cl ₂)	3mL
Iodine water (I ₂)	3mL
Mineral oil	36mL
0.1 M Potassium iodide solution (KI)	5mL
0.1 M Sodium bromide (NaBr)	5mL
0.1 M Sodium chloride (NaCl)	5mL
Labeling tape	1 roll

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S Marginson



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