

CHEMICAL ACTIVITY OF METALS

Lab Guidelines (1,2,3,4,5&6[attach guidelines],7,10,12)

Purposes:

To observe the reactivity of some common metals.

To understand the dynamics of a single replacement reaction, also known as an oxidation/reduction reaction.

To understand why some metals are more useful to us than others.

Introductory Information:

What is a single replacement reaction?

What is an oxidation/reduction reaction? What is going on with the electrons?

What does a more active metal do to a less active one?

What will I be looking for as evidence of chemical activity?

Materials:

Well plate, solutions of $\text{Cu}(\text{NO}_3)_2$, $\text{Mg}(\text{NO}_3)_2$, $\text{Zn}(\text{NO}_3)_2$, AgNO_3

4 pieces of each of the following metals: Cu, Mg, Zn, Ag.

Procedure:

In separate wells, add each metal to several drops of each solution and record your observations. Use the data table below to guide you. If the metals are not shiny, use steel wool to polish them before attempting the reactions.

HINT: One of the reactions is very faint, so make sure you allow 1 minute for a reaction to take place!!

Dispose of contents of well plate in garbage can. Then rinse well plate thoroughly with water.

Data:

Create a data table like the one below ON YOUR OWN PAPER:

Record your observations in this table. Please use the abbreviation "ppt" if a precipitate appears and note the color. If no reaction occurs, write "NR." HINT: One of the reactions is very faint.

	$\text{Cu}(\text{NO}_3)_2$	$\text{Mg}(\text{NO}_3)_2$	$\text{Zn}(\text{NO}_3)_2$	AgNO_3
Cu				
Mg				
Zn				
Ag				

Activity Of Metals Lab Answers

**V. K. Sally, Chhaya Srivastava, Goyal
Brothers Prakashan**



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