

Chemistry Exam review Answer Key

Topics:

Sample questions:

- Ch 10--Pg 338 #59-60, 71-79, 88, 93, 97, 100; pg 343 #2, 3, 5, 8-14
 - 59: find the moles of each atom from the formula, find the atomic mass of each element, multiply number of atoms by molar mass for each element and add together
 - 60 A. 98.0 g/mol, b. 76.0 g/mol, c. 100.1 g/mol, d. 132.1 g/mol, e. 89.0 g/mol
 - 71. 52.9% Al and 47.1% O
 - 72. 70% Fe and 30% O
 - 73. A. 5.9% H, 94.1% S; b. 22.6% N, 6.5% H, 19.4% C, 51.6% O; c. 41.7% Mg, 54.9% O, 3.4% H; d. 42.1% Na, 18.9% P, 39.0% O
 - 74. A. 3.33 g S, b. 5.65 g N, c. 40.6 g Mg, d. 152 g P
 - 75. D
 - 76. An empirical formula has the lowest whole-number ratio of elements; a. molecular, b. molecular, c. molecular and empirical, d. molecular and empirical
 - 77. A. molecular, b. molecular, c. empirical
 - 78. A
 - 79. A. C₃H₆O₃, b. Hg₂Cl₂
 - 88. A. CO, b. C₂O₂NH₅, c. Cl₂OC
 - 93. C₃H₆O₃
 - 97. A molecular formula is a whole-number multiple of its empirical formula
 - 100. A. C₉H₁₁O₂N, b. C₉H₁₁O₂N
 - 2. C
 - 3. A
 - 5. A
 - 8. E
 - 9. A
 - 10. D
 - 11. C
 - 12. C₃H₈O, 60.0 g/mol
 - 13. C₂H₅NO₂, 75.0 g/mol
 - 14. C₃H₆O₂, 74.0 g/mol
- Ch 11--Pg 377 #34, 36-37, 48-52, 54-55, 59; pg 381 #1, 3, 9-11
 - 34. A. reactants: sodium and water; products: hydrogen and sodium hydroxide; b. reactants: carbon dioxide and water; products: oxygen and glucose
 - 36. A formula is a unique identifier of a substance. A different formula would indicate a different substance, not the one that is taking part in the reaction you are trying to balance.
 - 37. A. $2\text{PbO}_2 (s) \rightarrow 2\text{PbO} (s) + \text{O}_2 (g)$, b. $2\text{Fe}(\text{OH})_3 (s) \rightarrow \text{Fe}_2\text{O}_3 (s) + 3\text{H}_2\text{O} (l)$, c. $(\text{NH}_4)_2\text{CO}_3 (s) \rightarrow 2\text{NH}_3 (g) + \text{H}_2\text{O} (g) + \text{CO}_2 (g)$, d. $\text{CaCl}_2 (aq) + \text{H}_2\text{SO}_4 (aq) \rightarrow \text{CaSO}_4 (s) + 2\text{HCl} (aq)$
 - 48. A. $\text{Cl}_2 (g) + 2\text{KI} (aq) \rightarrow \text{I}_2 (aq) + 2\text{KCl} (aq)$, b. $2\text{Fe} (s) + 6\text{HCl} (aq) \rightarrow 2\text{FeCl}_3 (aq) + 3\text{H}_2 (g)$, c. $\text{P}_4\text{O}_{10} (s) + 6\text{H}_2\text{O} (l) \rightarrow 4\text{H}_3\text{PO}_4 (aq)$
 - 49. A. $\text{Cl}_2 + 2\text{NaI} \rightarrow 2\text{NaCl} + \text{I}_2$, b. $2\text{NH}_3 \rightarrow \text{N}_2 + 3\text{H}_2$, c. $4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$
 - 50. A. $\text{ZnS} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{S} + \text{ZnSO}_4$, b. $\text{NaOH} + \text{HNO}_3 \rightarrow \text{H}_2\text{O} + \text{NaNO}_3$, c. $2\text{KF} + \text{Ca}(\text{NO}_3)_2 \rightarrow \text{CaF}_2 + 2\text{KNO}_3$
 - 51. A. $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{NaOH}$, b. $\text{H}_2 + \text{Br}_2 \rightarrow 2\text{HBr}$, c. $\text{Cl}_2\text{O}_7 + \text{H}_2\text{O} \rightarrow 2\text{HClO}_4$

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