

Calculating pH

How will it be used for the concentration differences here?

Why?

As buffers and other substances control pH in organisms, it is important to understand the ability of buffers to resist change. This property is called **buffering capacity** and is usually expressed as millimoles per litre of buffer that will change pH by one unit. Thus, a buffer with a buffering capacity of 10 mmol/L will resist 10 mmol/L of strong acid or base. The buffering capacity of a buffer is related to the concentration of the buffer components. A buffer with a concentration of 0.1 M will have a buffering capacity of 0.1 M.

Table 1 – Ion Concentrations for Acids and Bases

Acid	Substance	Acidic, Basic or Neutral?	$[H^+]$ (M)	$[OH^-]$ (M)	pH or pOH
1	0.1 M HCl (strong)	Acidic	1.0×10^{-1} M	1.0×10^{-14} M	1.0
2	0.001 M HCl (strong)	Acidic	1.0×10^{-3} M	1.0×10^{-11} M	3.0
3	0.00001 M HCl (strong)	Acidic	1.0×10^{-5} M	1.0×10^{-9} M	5.0
4	0.000001 M HCl (strong)	Acidic	1.0×10^{-6} M	1.0×10^{-8} M	6.0
5	0.0000001 M HCl (strong)	Acidic	1.0×10^{-7} M	1.0×10^{-7} M	7.0
6	0.00000001 M HCl (strong)	Acidic	1.0×10^{-8} M	1.0×10^{-6} M	8.0
7	0.000000001 M HCl (strong)	Acidic	1.0×10^{-9} M	1.0×10^{-5} M	9.0
8	0.0000000001 M HCl (strong)	Acidic	1.0×10^{-10} M	1.0×10^{-4} M	10.0
9	0.00000000001 M HCl (strong)	Acidic	1.0×10^{-11} M	1.0×10^{-3} M	11.0
10	0.000000000001 M HCl (strong)	Acidic	1.0×10^{-12} M	1.0×10^{-2} M	12.0
11	0.0000000000001 M HCl (strong)	Acidic	1.0×10^{-13} M	1.0×10^{-1} M	13.0

1. What are the $[H^+]$ and $[OH^-]$ in the 1 M HCl?

$[H^+] = 1.0 \times 10^{-1}$ M

2. In the 1 M HCl, what are the $[H^+]$ and $[OH^-]$ concentrations in the 1 M HCl?

$[OH^-] = 1.0 \times 10^{-14}$ M

3. What are the $[H^+]$ and $[OH^-]$ concentrations in the 0.001 M HCl?

$[H^+] = 1.0 \times 10^{-3}$ M

$[OH^-] = 1.0 \times 10^{-11}$ M

4. What are the $[H^+]$ and $[OH^-]$ concentrations in the 0.00001 M HCl?

$[H^+] = 1.0 \times 10^{-5}$ M

$[OH^-] = 1.0 \times 10^{-9}$ M

Answer Key Pogil Chemistry Calculating Ph

Mary Griffin



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