# Answers

#### Chapter 1

#### Exercises

- (a) CUCO -+CUC + CO
  - (b) 2Mg + O → 2MgO
  - (c) H SO + 2NaOH → Na SO + 2H O
  - (d) N2 " + \_\_ → 2NH3 "
  - (e) CH + 2O → CO + 2H O
- 2 (a)(2K + 2H O → 2KOH + H

QN OH + 30 → 200 + 3H O

COM 210 → 210CH

46003-x20/203 +

F(eD + 3C -+ 3CO + 2Fe

3 (a)(2C,H\_ + 13O, → 8CO, + 10H,O

4866. + 50. → 4NO + 6H,O

3(a) + 8HNO → 3CU(NO ) + 2NO + 4H O

6(MD + 2N H -+2N + 10H 0 + 0

460 + N + 150 - 800 + 14H 0 + 2N

- 4 (a) Sand and water: heterogeneous
- Sibilities: tretterogeneous.
- Sociar and water: homogeneous
- S(d) and iron filings: heterogeneous
- E(s)and and water; homogeneous
  - (f) Steel: homogeneous
  - 5 (a)(2KNO\_(b)) → 2KNO\_(b) + O\_(g)
- O(60 O<sub>2</sub>(s) + H.SO<sub>2</sub>(aq) -+ CaSO<sub>2</sub>(s) + CO<sub>2</sub>(q) + H Offi

264b) + 2H,O() -→ 2L/OH(ac) + H,(a)

P(dNO,Julisq) + 2NaO(laq) → PbOU(s) +

2NeNO,(eq)

(e) 2C,H,(g) + 9O,(g) → 6CO,(g) + 6H,O()

In the answers, exercises and text there is inconsistency about whether the water formed in combustion reactions is gas or liquid. At room temperature it will ultimately finish up as a liquid. ast shown here.

- X has diffused more quickly, so it must be a lighter gas, its particles have greater velocity than the particles of Y at the same temperature. (Note though that they will both have the same value for average kinetic energy.)
- 7 From the kinetic molecular theory we would expect a solid to be more dense than its liquid. and therefore that ice would sink in water.
- 8 Bubbles will be present through the volume of the liquid. A brown gas is visible above the brown liquid. As the two states are at the same temperature, the particles have the same average kinetic energy and are moving at the same speed. The inter-particle distances in the gas are significantly larger than those in the liquid.
- At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid chase.
- 10 Steam will condense on the skin, releasing energy as it forms liquid at the same temperature (E-D) on Figure 1.4). This is additional to the energy released when both the boiling water and the condensed steam cool on the surface of the skin.
- 12 AW 01 EX 12

2

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- 1. Understanding the eBook Answer Key For Pearson Custom Library For Chemistry
  - The Rise of Digital Reading Answer Key For Pearson Custom Library For Chemistry
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Answer Key For Pearson Custom Library For Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Key For Pearson Custom Library For Chemistry
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answer Key For Pearson Custom Library For Chemistry
  - Personalized Recommendations
  - Answer Key For Pearson Custom Library For Chemistry User Reviews and Ratings
  - Answer Key For Pearson Custom Library For Chemistry and Bestseller Lists

- 5. Accessing Answer Key For Pearson Custom Library For Chemistry Free and Paid eBooks
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  - Highlighting and Note-Taking Answer Key For Pearson Custom Library For Chemistry
  - Interactive Elements Answer Key For Pearson Custom Library For Chemistry
- 8. Staying Engaged with Answer Key For Pearson Custom Library For Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Key For Pearson Custom Library For Chemistry
- 9. Balancing eBooks and Physical Books Answer Key For Pearson Custom Library For Chemistry
  - Benefits of a Digital Library
  - $\circ$  Creating a Diverse Reading Collection Answer Key For Pearson Custom Library For Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key For Pearson Custom Library For Chemistry
  - Setting Reading Goals Answer Key For Pearson Custom Library For Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key For Pearson Custom Library For Chemistry
  - Fact-Checking eBook Content of Answer Key For Pearson Custom Library For Chemistry
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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