

15 • Chemical Kinetics**RATE LAWS**

1. Consider the reaction:
- $2 \text{NO(g)} + \text{O}_2\text{(g)} \rightarrow 2 \text{NO}_2\text{(g)}$

The following data were obtained from three experiments using the method of initial rates:

	Initial [NO] mol L ⁻¹	Initial [O ₂] mol L ⁻¹	Initial rate NO mol L ⁻¹ s ⁻¹
Experiment 1	0.010	0.010	2.5×10^{-5}
Experiment 2	0.020	0.010	1.0×10^{-4}
Experiment 3	0.010	0.020	5.0×10^{-5}

- Determine the order of the reaction for each reactant.
 - Write the rate equation for the reaction.
 - Calculate the rate constant.
 - Calculate the rate (in mol L⁻¹s⁻¹) at the instant when [NO] = 0.015 mol L⁻¹ and [O₂] = 0.0050 mol L⁻¹
 - At the instant when NO is reacting at the rate 1.0×10^{-4} mol L⁻¹s⁻¹, what is the rate at which O₂ is reactant and NO₂ is forming?
2. The reaction $2 \text{NO(g)} + 2 \text{H}_2\text{(g)} \rightarrow \text{N}_2\text{(g)} + 2 \text{H}_2\text{O(g)}$ was studied at 904 °C, and the data in the table were collected.

	Initial [NO] mol L ⁻¹	Initial [H ₂] mol L ⁻¹	Initial rate N ₂ mol L ⁻¹ s ⁻¹
Experiment 1	0.420	0.122	0.136
Experiment 2	0.210	0.122	0.0339
Experiment 3	0.210	0.244	0.0678
Experiment 4	0.105	0.488	0.0339

- Determine the order of the reaction for each reactant.
 - Write the rate equation for the reaction.
 - Calculate the rate constant at 904 °C.
 - Find the rate of appearance of N₂ at the instant when [NO] = 0.350 M and [H₂] = 0.205 M.
3. The reaction of 'butyl-bromide (CH₃)₃CBr with water is represented by the equation:



The following data were obtained from three experiments using the method of initial rates:

	Initial [(CH ₃) ₃ CBr] mol L ⁻¹	Initial [H ₂ O] mol L ⁻¹	Initial rate mol L ⁻¹ min ⁻¹
Experiment 1	5.0×10^{-2}	2.0×10^{-2}	2.0×10^{-6}
Experiment 2	5.0×10^{-2}	4.0×10^{-2}	2.0×10^{-6}
Experiment 3	1.0×10^{-1}	4.0×10^{-2}	4.0×10^{-6}

- What is the order with respect to (CH₃)₃CBr?
- What is the order with respect to H₂O?
- What is the overall order of the reaction?
- Write the rate equation.
- Calculate the rate constant, k, for the reaction.

Answer Key For Chemistry Reaction Rates

Rob Ritchie



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Table of Contents Answer Key For Chemistry Reaction Rates

1. Understanding the eBook Answer Key For Chemistry Reaction Rates
 - The Rise of Digital Reading Answer Key For Chemistry Reaction Rates
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Chemistry Reaction Rates
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Chemistry Reaction Rates
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Chemistry Reaction Rates
 - Personalized Recommendations
 - Answer Key For Chemistry Reaction Rates User Reviews and Ratings
 - Answer Key For Chemistry Reaction Rates and Bestseller Lists
5. Accessing Answer Key For Chemistry Reaction Rates Free and Paid eBooks
 - Answer Key For Chemistry Reaction Rates Public Domain eBooks
 - Answer Key For Chemistry Reaction Rates eBook Subscription Services
 - Answer Key For Chemistry Reaction Rates Budget-Friendly Options
6. Navigating Answer Key For Chemistry Reaction Rates eBook Formats

- ePub, PDF, MOBI, and More
- Answer Key For Chemistry Reaction Rates Compatibility with Devices
- Answer Key For Chemistry Reaction Rates Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For Chemistry Reaction Rates
 - Highlighting and Note-Taking Answer Key For Chemistry Reaction Rates
 - Interactive Elements Answer Key For Chemistry Reaction Rates
- 8. Staying Engaged with Answer Key For Chemistry Reaction Rates
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key For Chemistry Reaction Rates
- 9. Balancing eBooks and Physical Books Answer Key For Chemistry Reaction Rates
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For Chemistry Reaction Rates
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key For Chemistry Reaction Rates
 - Setting Reading Goals Answer Key For Chemistry Reaction Rates
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key For Chemistry Reaction Rates
 - Fact-Checking eBook Content of Answer Key For Chemistry Reaction Rates
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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