

THE MOLE AND AVOGADRO'S NUMBER

Name _____

One mole of a substance contains Avogadro's Number (6.02×10^{23}) of molecules.

How many molecules are in the quantities below?

1. 2.0 moles $x = 2.0 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{1.2 \times 10^{24} \text{ molecules}}$

2. 1.5 moles $x = 1.5 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{23} \text{ molecules}}$

3. 0.75 mole $x = 0.75 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{4.5 \times 10^{23} \text{ molecules}}$

4. 15 moles $x = 15 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{24} \text{ molecules}}$

5. 0.35 mole $x = 0.35 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{2.1 \times 10^{23} \text{ molecules}}$

How many moles are in the number of molecules below?

1. 6.02×10^{23} $x \text{ moles} = 6.02 \times 10^{23} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.0 \text{ mol}}$

2. 1.204×10^{24} $x \text{ moles} = 1.204 \times 10^{24} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.0 \text{ mol}}$

3. 1.5×10^{20} $x \text{ moles} = 1.5 \times 10^{20} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.5 \times 10^{-4} \text{ mol}}$ (0.00025 mol)

4. 3.4×10^{26} $x \text{ moles} = 3.4 \times 10^{26} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{565 \text{ mol}}$

5. 7.5×10^{19} $x \text{ moles} = 7.5 \times 10^{19} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.2 \times 10^{-4} \text{ mol}}$

Answer Key Chemistry If8766 Moles And Mass

John Lambert



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CONTENTS

- 1 Writing the Formulae of Ionic Compounds
- 2 Formula Masses
- 3 Mass and Percent Compositions
- 4 The Mole and Mass
- 5 The Mole and Number of Particles
- 6 The Mole and Concentration
- 7 The Mole and Molar Volume
- 8 Empirical and Molecular Formulae
- 9 Chemical Equations
- 10 Calculations Based on Chemical Equations
- 11 Limiting Reagent
- 12 Gas Volume Volume Stoichiometry
- 13 Percent Yield

Answers to Practice Problems Appendices

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1. Understanding the eBook Answer Key Chemistry If8766 Moles And Mass
 - The Rise of Digital Reading Answer Key Chemistry If8766 Moles And Mass
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Chemistry If8766 Moles And Mass
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Chemistry If8766 Moles And Mass
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Chemistry If8766 Moles And Mass
 - Personalized Recommendations
 - Answer Key Chemistry If8766 Moles And Mass User Reviews and Ratings
 - Answer Key Chemistry If8766 Moles And Mass and Bestseller Lists
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 - Answer Key Chemistry If8766 Moles And Mass Enhanced eBook Features
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 - Highlighting and Note-Taking Answer Key Chemistry If8766 Moles And Mass
 - Interactive Elements Answer Key Chemistry If8766 Moles And Mass
8. Staying Engaged with Answer Key Chemistry If8766 Moles And Mass
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Chemistry If8766 Moles And Mass
9. Balancing eBooks and Physical Books Answer Key Chemistry If8766 Moles And Mass
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Chemistry If8766 Moles And Mass
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Chemistry If8766 Moles And Mass
- Setting Reading Goals Answer Key Chemistry If8766 Moles And Mass
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Chemistry If8766 Moles And Mass
- Fact-Checking eBook Content of Answer Key Chemistry If8766 Moles And Mass
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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