Chemistry Student Edition - Basic Answer Key

Chapter 19: Chemical Equilibrium

The Nature of Chemical Equilibrium

Questions

- At chemical equilibrium, what is the relationship between the rate of the forward reaction and the rate of the reverse reaction?
- 2. What does the law of mass action say about any reaction at equilibrium?
- 3. Why are solids and liquids left out of the equilibrium constant expression?
- Explain why the equilibrium constant expression as derived from concentrations can be converted to the equilibrium constant expression as derived from partial pressures.
- Construct balanced chemical equations and write the equilibrium constant expression (either K_a or K_a) for each of the following processes:
 - a) Sodium metal reacts vigorously with oxygen gas to produce sodium oxide.
 - b) Aluminum oxide decomposes to form aluminum metal and oxygen gas.
 - Nitrogen gas and oxygen gas react to form nitrogen monoxide gas.
- Give an example of a chemical equilibrium that exists in our daily lives.
- Write a K_{**} expression for each of the following processes.
 - Solid calcium carbonate is heated and decomposes into solid calcium oxide and carbon dioxide gas.
 - Hydrogen gas (H₂) and solid iodine (I₂) are combined to form hydrogen iodide gas (HI).

Answers

5.

- 1. At equilibrium, the rate of the forward reaction equals the rate of the reverse reaction
- The ratio of product concentrations to reactant concentrations (along with the appropriate exponents) is equal to a constant.
- Solids and liquids have essentially constant densities, which means that writing them is not very useful.
- 4. Partial pressures are related to concentrations by a factor of RT.

$$\begin{split} &2\mathrm{Na}_{00}+\Diamond_{200} \stackrel{\longleftarrow}{\longleftarrow} \mathrm{\acute{E}}\ 2\mathrm{Na}_{2}\Diamond_{00}\\ &\mathrm{K}_{c}=\frac{[\mathrm{Na}_{2}\varnothing_{00}]^{2}}{[\mathrm{Na}_{00}]^{2}[\Diamond_{200}]} \end{split}$$

a) heterogeneous equilibria

$$2Al_2 \bigcirc_{300} \longrightarrow 4Al_{00} + 3\bigcirc_{300}$$

$$K_c = \frac{[AI'_{00}]^4 [\bigcirc_{200}]^3}{[AI_2 \varnothing_{80}]^2}$$

b) heterogeneous equilibria

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- 1. Understanding the eBook Answer Key Ch 19 Chemistry
 - The Rise of Digital Reading Answer Key Ch 19 Chemistry
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Answer Key Ch 19 Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Ch 19 Chemistry
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answer Key Ch 19 Chemistry
 - Personalized Recommendations
 - Answer Key Ch 19 Chemistry User Reviews and Ratings
 - Answer Key Ch 19 Chemistry and Bestseller Lists
- 5. Accessing Answer Key Ch 19 Chemistry Free and Paid eBooks

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- Answer Key Ch 19 Chemistry eBook Subscription Services
- Answer Key Ch 19 Chemistry Budget-Friendly Options
- 6. Navigating Answer Key Ch 19 Chemistry eBook Formats
 - o ePub, PDF, MOBI, and More
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 - Answer Key Ch 19 Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Ch 19 Chemistry
 - Highlighting and Note-Taking Answer Key Ch 19 Chemistry
 - Interactive Elements Answer Key Ch 19 Chemistry
- 8. Staying Engaged with Answer Key Ch 19 Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Ch 19 Chemistry
- 9. Balancing eBooks and Physical Books Answer Key Ch 19 Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Ch 19 Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key Ch 19 Chemistry
 - Setting Reading Goals Answer Key Ch 19 Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key Ch 19 Chemistry
 - Fact-Checking eBook Content of Answer Key Ch 19 Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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