

Limiting Reactants Practice



a) What is the limiting reactant if 0.866 mol of NO is mixed with 0.503 mol of O₂ to create NO₂?

$$0.866 \text{ mol NO} \times \frac{2 \text{ mol NO}_2}{2 \text{ mol NO}} = 0.866 \text{ mol NO}_2$$

limiting reactant product made

$$0.503 \text{ mol O}_2 \times \frac{2 \text{ mol NO}_2}{1 \text{ mol O}_2} = 1.006 \text{ mol NO}_2$$

b) Calculate the number of moles of NO₂ produced. (Answer should already be solved from the previous questions)

0.866 mol NO₂



a) If 7.40 g of O₃ reacts with 0.670 g of NO, which compound will be the limiting reactant?

O₃
O: 3 × 15.999 = 47.997 g/mol
⇒ 1 mol O₃ = 47.997 g O₃

$$7.40 \text{ g O}_3 \times \frac{1 \text{ mol O}_3}{47.997 \text{ g O}_3} \times \frac{1 \text{ mol NO}_2}{1 \text{ mol O}_3} \times \frac{46.005 \text{ g NO}_2}{1 \text{ mol NO}_2} = 7.093 \text{ g NO}_2$$

NO
N: 1 × 14.007 = 14.007
O: 1 × 15.999 = 15.999
⇒ 1 mol NO = 30.006 g/mol

$$0.670 \text{ g NO} \times \frac{1 \text{ mol NO}}{30.006 \text{ g NO}} \times \frac{1 \text{ mol NO}_2}{1 \text{ mol NO}} \times \frac{46.005 \text{ g NO}_2}{1 \text{ mol NO}_2} = 1.027 \text{ g NO}_2$$

limiting reactant product made

b) How many grams of NO₂ will be produced? (Answer should already be solved from the previous questions)

1.027 g NO₂

c) Calculate mass, in grams, of the excess reactant remaining at the end of the reaction. How many moles of excess reactant would this be?

$$0.670 \text{ g NO} \times \frac{1 \text{ mol NO}}{30.006 \text{ g NO}} \times \frac{1 \text{ mol O}_3}{1 \text{ mol NO}} \times \frac{47.997 \text{ g O}_3}{1 \text{ mol O}_3} = 1.072 \text{ g O}_3 \text{ used!}$$

$$\underbrace{7.40 \text{ g O}_3}_{\text{original}} - \underbrace{1.072 \text{ g O}_3}_{\text{used}} = \underline{6.328 \text{ g O}_3 \text{ leftover}}$$

Answer Key To Pogil 33 Limiting Reactants

Lifelique

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

Answer Key To Pogil 33 Limiting Reactants:

Limiting Reactant and Percent Yield, 2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

The Enigmatic Realm of **Answer Key To Pogil 33 Limiting Reactants**: Unleashing the Language is Inner Magic

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