

Section 18.4 Neutralization

In your textbook, read about neutralization and titration.

For each item in Column A, write the letter of the matching item in Column B.

Column A

- _____ 1. A chemical dye that changes color based on the pH of a solution
- _____ 2. A method for using a neutralization reaction to determine the concentration of a solution
- _____ 3. A reaction in which an acid and a base react to produce a salt and water
- _____ 4. A solution of known concentration
- _____ 5. An ionic product of an acid-base reaction
- _____ 6. The point in a titration in which an indicator changes color
- _____ 7. The stoichiometric point of a titration

Column B

- a. acid-base indicator
- b. end point
- c. equivalence point
- d. neutralization
- e. salt
- f. standard solution
- g. titration

Complete the following table, indicating the formula and name of the salt formed by a neutralization reaction between the listed acid and base.

Acid	Base	Salt formula	Salt name
8. HCl	KOH	KCl	potassium chloride
9. H_2SO_4	$Mg(OH)_2$		
10. H_3PO_4	NaOH		
11. HNO_3	$Fe(OH)_3$		
12. H_3PO_4	$Ca(OH)_2$		

In the space at the left, write 1 through 4 to show the correct sequence of the steps in performing a titration using a pH meter. Then, write 5 through 8 to sequence the steps used to calculate the concentration of the unknown solution.

Sequence of Steps

- _____ 13. Continue adding the standard solution to the solution of unknown concentration until the equivalence point is reached.
- _____ 14. Fill a buret with the standard solution.
- _____ 15. Start adding the standard solution slowly, with mixing, to the solution of unknown concentration, reading the pH at regular intervals.

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RS Peters



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Chemistry Steven S. Zumdahl, Susan A. Zumdahl, 2000 Zumdahl's conceptual approach to problem solving early and thorough treatment of reactions and integration of descriptive chemistry and chemical principles throughout are key elements to the success of this bestseller By encouraging students to understand qualitative concepts before considering quantitative problems Chemistry helps students build a solid understanding of topics and avoid common misconceptions The text also includes new Chemical Impact boxes that present everyday applications of chemistry plus over 200 redesigned illustrations with greater clarity and emphasis on molecular level illustrations **Resources in Education**, 1984 Serves as an index to Eric reports microform *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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