#### ACIDS, BASES, & THE pH SCALE

Name: Answer Key
Date:
Pd:

#### PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
<ul> <li>Sour taste, leaves burning sensation</li> </ul>	<ul> <li>Sharp <u>bitter</u> taste, feels <u>slippery</u> to touch</li> </ul>
<ul> <li>Increase hydrogen ions (H+) in solution</li> </ul>	<ul> <li>Decrease hydrogen ions (H+) in solution (Increase hydroxide ions (OH-) in solution)</li> </ul>
• Example: HCI H'+CI	Example: NaOH Na' + OH
<ul> <li>Turns <u>litmus</u> (pH) paper to <u>red</u></li> </ul>	<ul> <li>Turns litmus (pH) paper to blue</li> </ul>

#### PART 2: THE pH SCALE

- The pH Scale is a measurement of how acidic or how basic a solution is.
- The scale goes from 0 to 14.
- . Compounds are acidic if they have a pH lower than Z.
- Compounds with a pH <u>higher</u> than <u>7</u> are <u>basic</u> or <u>alkaline</u>.
- Compounds with a pH of 7, are neutral.

0 7 8 9 10 11 12 13

Increased concentration of H ions
Decreased concentration of OH ions

Increased concentration of OH ions

pH below 7 = acids

pH at  $7 = \frac{neutral}{pure water}$ 

pH above 7 = <u>bases</u>

Decreased concentration of H' ions

#### PART 3: BUFFERS

- A <u>buffer</u> is a <u>weak acid</u> or <u>base</u> that <u>prevents</u> sudden <u>changes</u> in pH.
  - Maintains homeostasis
- A buffer works by binding to H+ ions releasing H+ ions when the solution pH changes.
- Buffers bind to H+ ions when the concentration of H= ions in solution increases.
- NaOH Na + OH (The OH- will bind to H+ ions to form H;O neutralization)
- Buffers release H+ ions when the concentration of H= ions in solution decreases.
   HCI H' + CI

# **Answer Key Acids Bases**

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