

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **increases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Answer Key To Periodic Trends

Dr. Ashokan Kannarath



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Table of Contents Answer Key To Periodic Trends

1. Understanding the eBook Answer Key To Periodic Trends
 - The Rise of Digital Reading Answer Key To Periodic Trends
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Periodic Trends
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To Periodic Trends
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Periodic Trends
 - Personalized Recommendations
 - Answer Key To Periodic Trends User Reviews and Ratings
 - Answer Key To Periodic Trends and Bestseller Lists
5. Accessing Answer Key To Periodic Trends Free and Paid eBooks

- Answer Key To Periodic Trends Public Domain eBooks
- Answer Key To Periodic Trends eBook Subscription Services
- Answer Key To Periodic Trends Budget-Friendly Options
- 6. Navigating Answer Key To Periodic Trends eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key To Periodic Trends Compatibility with Devices
 - Answer Key To Periodic Trends Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To Periodic Trends
 - Highlighting and Note-Taking Answer Key To Periodic Trends
 - Interactive Elements Answer Key To Periodic Trends
- 8. Staying Engaged with Answer Key To Periodic Trends
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To Periodic Trends
- 9. Balancing eBooks and Physical Books Answer Key To Periodic Trends
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To Periodic Trends
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key To Periodic Trends
 - Setting Reading Goals Answer Key To Periodic Trends
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key To Periodic Trends
 - Fact-Checking eBook Content of Answer Key To Periodic Trends
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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