

Chapter 8 Balancing Equations Set # 3

Using the equations from Practice set #1 and #2 perform the following:

1. Balance the equation
2. Classify the balanced equation using the following choices:
 - Synthesis
 - Decomposition
 - Single Displacement (Cation or Anion)
 - Combustion (Hydrocarbon or Carbohydrate)
 - Double Displacement / Double Displacement (Acid/Base)
 - Other

Set #1

1. $\text{HF (g)} \longrightarrow \text{H}_2 \text{ (g)} + \text{F}_2 \text{ (g)}$ _____
2. $\text{Al (s)} + \text{I}_2 \text{ (s)} \longrightarrow \text{AlI}_3 \text{ (aq)}$ _____
3. $\text{FeO (s)} + \text{O}_2 \text{ (g)} \longrightarrow \text{Fe}_2\text{O}_3 \text{ (s)}$ _____
4. $\text{C}_4\text{H}_{10} \text{ (s)} + \text{O}_2 \text{ (g)} \longrightarrow \text{CO}_2 \text{ (g)} + \text{H}_2\text{O (l)}$ _____
5. $\text{Na (s)} + \text{MgI}_2 \text{ (aq)} \longrightarrow \text{NaI (aq)} + \text{Mg (s)}$ _____
6. $\text{AgNO}_3 \text{ (s)} + \text{Na}_2\text{S (aq)} \longrightarrow \text{Ag}_2\text{S (s)} + \text{NaNO}_3 \text{ (aq)}$ _____
7. $\text{K}_2\text{SO}_4 \text{ (aq)} + \text{Mg(NO}_3)_2 \text{ (aq)} \longrightarrow \text{MgSO}_4 \text{ (ppt)} + \text{KNO}_3 \text{ (aq)}$ _____
8. $\text{Al}_2(\text{CO}_3)_3 \text{ (s)} \longrightarrow \text{Al}_2\text{O}_3 \text{ (s)} + \text{CO}_2 \text{ (aq)}$ _____
9. $\text{Li (s)} + \text{AuCl}_3 \text{ (aq)} \longrightarrow \text{LiCl (aq)} + \text{Au (s)}$ _____
10. $\text{Li}_2\text{CrO}_4 \text{ (aq)} + \text{BaCl}_2 \text{ (aq)} \longrightarrow \text{LiCl (aq)} + \text{BaCrO}_4 \text{ (ppt)}$ _____
11. $\text{C}_{12}\text{H}_{22}\text{O}_{11} \text{ (s)} + \text{O}_2 \text{ (g)} \longrightarrow \text{CO}_2 \text{ (g)} + \text{H}_2\text{O (l)}$ _____
12. $\text{Ca(OH)}_2 \text{ (aq)} + \text{H}_2\text{SO}_4 \text{ (aq)} \longrightarrow \text{CaSO}_4 \text{ (aq)} + \text{H}_2\text{O (l)}$ _____

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Table of Contents Answer Key Instructional Fair Inc Balance Equation

1. Understanding the eBook Answer Key Instructional Fair Inc Balance Equation
 - The Rise of Digital Reading Answer Key Instructional Fair Inc Balance Equation
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Instructional Fair Inc Balance Equation
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Instructional Fair Inc Balance Equation
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Instructional Fair Inc Balance Equation
 - Personalized Recommendations
 - Answer Key Instructional Fair Inc Balance Equation User Reviews and Ratings
 - Answer Key Instructional Fair Inc Balance Equation and Bestseller Lists

5. Accessing Answer Key Instructional Fair Inc Balance Equation Free and Paid eBooks
 - Answer Key Instructional Fair Inc Balance Equation Public Domain eBooks
 - Answer Key Instructional Fair Inc Balance Equation eBook Subscription Services
 - Answer Key Instructional Fair Inc Balance Equation Budget-Friendly Options
6. Navigating Answer Key Instructional Fair Inc Balance Equation eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key Instructional Fair Inc Balance Equation Compatibility with Devices
 - Answer Key Instructional Fair Inc Balance Equation Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Instructional Fair Inc Balance Equation
 - Highlighting and Note-Taking Answer Key Instructional Fair Inc Balance Equation
 - Interactive Elements Answer Key Instructional Fair Inc Balance Equation
8. Staying Engaged with Answer Key Instructional Fair Inc Balance Equation
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Instructional Fair Inc Balance Equation
9. Balancing eBooks and Physical Books Answer Key Instructional Fair Inc Balance Equation
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Instructional Fair Inc Balance Equation
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Instructional Fair Inc Balance Equation
 - Setting Reading Goals Answer Key Instructional Fair Inc Balance Equation
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Instructional Fair Inc Balance Equation
 - Fact-Checking eBook Content of Answer Key Instructional Fair Inc Balance Equation
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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