

PAPER 2

2 hours

ESSAY

[100 MARKS]

Answer **four** questions in **all**: question 1 in Section A and three questions from Section B.

All questions carry **equal** marks.

No marks will be awarded for answering questions **not** specified for your own country.

Credit will be given for clarity of expression and orderly presentation of materials.

SECTION A

Answer **all** the questions in this section.

1.
 - (a)
 - (i) Define an *acid* according to the Lewis concept.
 - (ii) Give **one** example of a Lewis acid. [3 marks]
 - (b) Explain salting out in soap preparation. [2 marks]
 - (c) State the reagent and condition necessary for the following conversion:
 $\text{H} - \text{C} \equiv \text{C} - \text{H} \rightarrow \text{Ag} - \text{C} \equiv \text{C} - \text{Ag}$ [2 marks]
 - (d) What is the percentage abundance of an isotope? [2 marks]
 - (e)
 - (i) Why does the element with atomic number 18 **not** have an oxide?
 - (ii) Explain why chlorine (II) oxide has a low melting point. [3 marks]
 - (f) Describe a test to distinguish between concentrated HNO_3 and concentrated H_2SO_4 . [3 marks]
 - (g) State **two** differences between an electrochemical cell and an electrolytic cell. [3 marks]
 - (h) How does the trend in ionization energy affect the reactivity of group 1 elements? [3 marks]
 - (i) Define the term *molecular formula*. [2 marks]
 - (j)
 - (i) State which of the gases Cl_2 and NH_3 would deviate **more** from ideal behaviour.
 - (ii) Give reasons for the answer stated in 1(j) (i) [3 marks]

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