

CHEMISTRY SUBJECT TEST 3—*Continued*

PLEASE GO TO THE SPECIAL SECTION LABELED CHEMISTRY AT THE LOWER RIGHT-HAND CORNER OF THE ANSWER SHEET YOU ARE WORKING ON AND ANSWER QUESTIONS 101–116 ACCORDING TO THE FOLLOWING DIRECTIONS.

Part B

Directions: Each question below consists of two statements, I in the left-hand column and II in the right-hand column. For each question, determine whether statement I is true or false and whether statement II is true or false, and fill in the corresponding T or F ovals on your answer sheet. Fill in oval CE only if statement II is a correct explanation of statement I.

EXAMPLES:															
I		II													
EX 1. H_2SO_4 is a strong acid	BECAUSE	H_2SO_4 contains sulfur.													
EX 2. An atom of oxygen is electrically neutral	BECAUSE	an oxygen atom contains an equal number of protons and electrons.													
<div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;"> <p>SAMPLE ANSWERS</p> <table border="1" style="margin: auto;"> <tr> <th></th> <th>I</th> <th>II</th> <th>CE</th> </tr> <tr> <td>EX 1</td> <td style="text-align: center;"><input checked="" type="radio"/></td> <td style="text-align: center;"><input type="radio"/></td> <td style="text-align: center;"><input type="radio"/></td> </tr> <tr> <td>EX 2</td> <td style="text-align: center;"><input type="radio"/></td> <td style="text-align: center;"><input checked="" type="radio"/></td> <td style="text-align: center;"><input type="radio"/></td> </tr> </table> </div> </div>					I	II	CE	EX 1	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>	EX 2	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>
	I	II	CE												
EX 1	<input checked="" type="radio"/>	<input type="radio"/>	<input type="radio"/>												
EX 2	<input type="radio"/>	<input checked="" type="radio"/>	<input type="radio"/>												

I		II	
101. Transition metal compounds are often colored	BECAUSE	they frequently possess partially filled <i>d</i> orbitals.	
102. Chemical reactions slow down with lower temperature	BECAUSE	the energy barrier for the formation of products decreases with decreasing temperature.	
103. Exothermic reactions absorb heat	BECAUSE	breaking covalent bonds always requires energy.	
104. The solubility of gases in liquids does not depend upon pressure	BECAUSE	the vapor pressure of a substance is independent of external pressure.	
105. MgO has a high melting point	BECAUSE	highly charged ions result in strong ionic forces and high lattice energies.	
106. The ground state electron configuration orbitals of elemental Cu is $[\text{Ar}] 4s^1 3d^{10}$	BECAUSE	completely half-filled and filled <i>d</i> bestow special electronic stabilization.	
107. Isotopes of a particular element have nearly identical chemical behavior	BECAUSE	they have identical electron configurations.	
108. In an electrochemical cell, the electrode that is the site of reduction is called the anode	BECAUSE	oxidation always occurs at the cathode.	

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