

34. The maximum number of covalent bonds formed by nitrogen is

A. 1.
B. 2.
C. 3.
D. 4.

35. The IUPAC name of the compound $\text{CH}_3\text{CH}(\text{CH}_3)\text{CHCH}_3$ is

A. 2 - methyl but - 1 - ene.
B. 2 - methyl but - 2 - ene.
C. 3 - methyl but - 1 - ene.
D. 3 - methyl but - 2 - ene.

36. Ionization energy increases across the period in the periodic table because

A. atomic number increases.
B. effective nuclear charge increases.
C. mass number decreases.
D. screening effect decreases.

37. Which of the following properties indicate that an element is a metal. It

I. reacts with oxygen to form an acidic oxide
II. forms ionic chlorides.
III. has variable oxidation states.
IV. displaces hydrogen from dilute HCl.

A. I and III only
B. I and II only
C. II and IV only
D. I, II, III and IV

38. The electron configuration of carbon atom in its excited state is

A. $1s^2 2s^2 2p_x^1 2p_y^0$
B. $1s^2 2s^2 2p_x^1 2p_y^1$
C. $1s^2 2s^1 2p_x^1 2p_y^1 2p_z^1$
D. $1s^1 2s^2 2p_x^1 2p_y^1 2p_z^1$

39. An oxide has the following properties. It

I. is a white powder.
II. reacts with HCl.
III. reacts with NaOH.
IV. is insoluble in water.

The oxide is

A. alkaline.
B. amphoteric.
C. acidic.
D. neutral.

40. Which of the following statements about atoms of a metal is correct? They

A. readily accept electrons.
B. are soft.
C. are held together by covalent bond.
D. are held together by a sea of electron cloud.

41. Which of the following sketches is a graphical representation of Boyle's law?



42. The atom with the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^4$ is in

A. period 4, p- block.
B. period 3, p- block.
C. period 4, d- block.
D. period 3, d- block.

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