



A solution of 0.100 M HCl and a solution of 0.100 M NaOH are prepared. A 40.0 mL sample of one of the solutions is added to a beaker and then titrated with the other solution. A pH electrode is used to obtain the data that are plotted in the titration curve shown above.

- (a) Identify the solution that was initially added to the beaker. Explain your reasoning.

1/1

The solution in the beaker was the 0.100 M HCl because the initial pH was 1 (the pH of 0.100 M HCl).

One point is earned for the correct identification with rationale.

- (b) On the titration curve above, circle the point that corresponds to the equivalence point.

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