M.S. By Research (Thory)
Chomistry: People I Physical Chomistry P2-wpg-Teb-13KL-156 Con. 1359-13. [Total Marks: 100 (3 Hours) 18th April. 2013 N.B. : (1) All questions are compulsory. (2) Figures to the right indicate maximum marks. (3) Use of non-programmable scientific calculator is allowed. Useful constants: $c = 2.998 \times 10^8 \text{ms}^{-1}$ $h = 6.626 \times 10^{-34} \text{ Js}$ $R = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$ m. = 9-110 x 10-31 kg $k = 1.3811 \times 10^{-23} JK^{-1}$ $N = 6.022 \times 10^{23} \text{ mol}^{-1}$ $e = 1.602 \times 10^{-19} C$ $1 J = 6.24 \times 10^{18} \text{ eV}$ $1 \text{ eV} = 8.06 \times 10^3 \text{ cm}^{-1}$ 20 Attempt any five of the following: (a) What is fugacity? How is it evaluated for a real gas? (b) State the expression for Boltzman distribution function. Explain the terms involved in it. (c) Explain giving reasons the important applications of fuel-cells. (d) State Debye-Huckel Onsager equation. How its validity can be checked graphically? (e) Determine the degree of degeneracy for energy levels $\frac{11h^2}{2}$ for a particle in cubical box. (f) What is state function in quantum mechanics? Give its characteristics. (g) Explain :-(i) Potential energy surfaces Oscillating reactions. (44) (h) Predict the effect of ionic strength on rate constant of the following reactions:

 $C_{12} H_{22} O_{11} + H^+ \rightarrow Products$

2Cu2+ + 41- → 2Cu1 + I2 Cii)

(iii) $S_2O_8^{2-} + 21^- \rightarrow 2SO_4^{2-} + 1$,

(iv) $\left[\text{Pt Cl}_4 \right]^{2-} + \text{OH}^- \rightarrow \text{Products}$

2. (a) What is partition function? Derive an expression for rotational partition function 7 of a diatomic molecule. SHO

(a) Explain the terms :— (i) Joule-Thomson coefficient (ii) Inversion temperature.

(b) (i) Calculate the molar residual entropy of a crystal in which molecules can adopt five orientations of equal energy at absolute zero.

(ii) Calculate the value of molar residual entropy for NO molecules at absolute zero.

7

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L. H. M. Chung

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- 1. Understanding the eBook Chemistry Question Papers Tndte April
 - The Rise of Digital Reading Chemistry Question Papers Tndte April
 - $\circ\,$ Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry Question Papers Tndte April
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Question Papers Tndte April
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry Question Papers Tndte April
 - Personalized Recommendations
 - Chemistry Question Papers Tndte April User Reviews and Ratings
 - Chemistry Question Papers Tndte April and Bestseller Lists
- 5. Accessing Chemistry Question Papers Tndte April Free and Paid eBooks
 - Chemistry Question Papers Tndte April Public Domain eBooks
 - Chemistry Question Papers Tndte April eBook Subscription Services

- Chemistry Question Papers Tndte April Budget-Friendly Options
- 6. Navigating Chemistry Question Papers Tndte April eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chemistry Question Papers Tndte April Compatibility with Devices
 - Chemistry Question Papers Tndte April Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Question Papers Tndte April
 - Highlighting and Note-Taking Chemistry Question Papers Tndte April
 - Interactive Elements Chemistry Question Papers Tndte April
- 8. Staying Engaged with Chemistry Question Papers Tndte April
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Question Papers Tndte April
- 9. Balancing eBooks and Physical Books Chemistry Question Papers Tndte April
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Question Papers Tndte April
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Question Papers Tndte April
 - Setting Reading Goals Chemistry Question Papers Tndte April
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Question Papers Tndte April
 - Fact-Checking eBook Content of Chemistry Question Papers Tndte April
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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