

- (xiv) This has five significant figure:
 • 0.00215 • 21500 • 21.5210 • 21.521
- (xv) The melting points of those substances which expand on melting increase when the pressure is:
 • decreased • increased • kept unchanged • atmospheric
- (xvi) While considering the fule of significant figures the simplified result of $3.1 + 14.367$ is:
 • 17.467 • 17.47 • 17.50 • 17.5
- (xvii) The atmospheric pressures recorded in different places at the same time are given below:

Nathiagali	Hunza	Murree	Gilgit
700 torr	650 torr	710 torr	600 torr

Water will boil first in:

- Nathiagali • Hunza • Murree • Gilgit

CHEMISTRY

2017

Time: 2 Hours 40 Minutes

Marks: 68

SECTION 'B' (SHORT-ANSWER QUESTIONS)

NOTE: Answer any Ten part questions. (40)

- 2.(i) Define the following:
 (1) Limiting reactant (2) Internal energy
 (3) Mole (4) Threshold energy
- (ii) An organic compound contains 40% C and 6.66% H by mass. Its molecular mass is double that of its empirical formula mass. Find its Molecular formula.
- (iii) $\text{CaCO}_3 \xrightarrow{\text{heat}} \text{CaO} + \text{CO}_2$ CaCO_3 is often used to generate CO_2 gas in industry. If 200g of CaCO_3 is strongly heated, what volume of CO_2 gas will be obtained at 30°C and 1200 torr pressure?
- (iv) Give the reasons for any four of the following:
 (a) Glycerine is distilled at reduced pressure.
 (b) Milk sours more rapidly in summer than in winter.
 (c) Compounds having Hydrogen bond generally have high boiling points.
 (d) The shorter the bond length, the stronger will be the bond.
 (e) Elements give line spectrum.
- (v) A gaseous mixture of H_2 and O_2 at 60°C and total pressure of 2.5 atmospheres contains 75 mole % H_2 . What are the partial pressures of both the gases in the mixture.
- (vi) State the relevant gas laws which explain the following observations of gases:
 • When a balloon filled with air is kept in cold place at constant pressure, its volume is reduced.
 • When the pressure on a gas kept in an airlight cylinder at constant temperature is doubled, its volume becomes half.

Chemistry Board Paper Xi

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