

1 – REVIEW - Matter and Measurement

1. Identify the following changes as physical or chemical changes:
 - a. Baking soda reacts with vinegar to produce carbon dioxide. **CHEMICAL**
 - b. The copper sheath on the Statue of Liberty turns green. **CHEMICAL**
 - c. Addition of salt melts ice on the highway. **PHYSICAL**
 - d. Steam condenses on the windowpane. **PHYSICAL**
 - e. Epoxy resin cures and hardens. **CHEMICAL**
 - f. Sugar dissolves in a cup of coffee. **PHYSICAL**
 - g. Natural gas burns in a furnace. **CHEMICAL**
2.
 - a. Calculate the density of lead if a 10 kg block has a volume of 885 cm³.
 $D = M/V = 10 \text{ kg} / 885 \text{ cm}^3 = 0.011 \text{ kg/cm}^3$
 - b. What is the volume of a 100 g bar of aluminum if its density is 2.70 g·cm⁻³?
 $V = M/D = 100 \text{ g} / 2.70 \text{ g·cm}^{-3} = 37 \text{ cm}^3$
 - c. Calculate the mass of 100 cm³ of uranium (density 19.07 g·cm⁻³).
 $M = DV = (19.07 \text{ g·cm}^{-3})(100 \text{ cm}^3) = 1907 \text{ g}$
4. Convert:
 - c. 25°C to K **$K = ^\circ\text{C} + 273.15 = 298.15 \text{ K}$**
6. Which of the following physical properties are extensive?
 - a. heat of fusion
 - b. melting point
 - c. **color**
 - d. **viscosity**
 - e. conductivity
 - f. density
7. Write the names of the following elements:
 - a. N **nitrogen**
 - b. Ca **calcium**
 - c. K **potassium**
 - d. P **phosphorous**
 - e. V **vanadium**
8. Write the symbols for the following elements:
 - a. silicon **Si**
 - b. chlorine **Cl**
 - c. iron **Fe**
 - d. sodium **Na**
 - e. silver **Ag**
 - f. sulfur **S**
9. Convert:
 - a. 1342 mL into L
 $1342 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 1.342 \text{ L}$
 - b. 3.26 x 10⁻⁶ km into mm
 $3.26\text{E-}6 \text{ km} \times \frac{1,000,000 \text{ mm}}{1 \text{ km}} = 3.26 \text{ mm}$

4 Nuclear Chemistry Answer South Pasadena

Annelies Wilder-Smith



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Analytical Chemistry in Nuclear Reactor Technology C. D. Susano,1959 *Nuclear Science Abstracts* ,1975

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