

Balancing Chemical Reactions

Balance all of the chemical reactions below. All compounds are written correctly.

After balancing the reaction, determine the type(s) of chemical reaction(s).

TYPE(S) of RXN

- | | |
|---|-------------------|
| 1. <u>2</u> $\text{NH}_3 \rightarrow$ _____ $\text{N}_2 +$ <u>3</u> H_2 | <u>Decomp</u> |
| 2. <u>4</u> $\text{Al} +$ <u>3</u> $\text{O}_2 \rightarrow$ <u>2</u> Al_2O_3 | <u>Syn / Comb</u> |
| 3. _____ $\text{N}_2\text{H}_4 +$ _____ $\text{O}_2 \rightarrow$ _____ $\text{N}_2 +$ <u>2</u> H_2O | <u>S.R. (?)</u> |
| 4. <u>2</u> $\text{Al} +$ <u>3</u> $\text{H}_2\text{SO}_4 \rightarrow$ <u>3</u> $\text{H}_2 +$ _____ $\text{Al}_2(\text{SO}_4)_3$ | <u>S.R.</u> |
| 5. _____ $\text{P}_2\text{O}_5 +$ <u>3</u> $\text{H}_2\text{O} \rightarrow$ <u>2</u> H_3PO_4 | <u>Syn</u> |
| 6. <u>2</u> $\text{C}_2\text{H}_2 +$ <u>5</u> $\text{O}_2 \rightarrow$ <u>4</u> $\text{CO}_2 +$ <u>2</u> H_2O | <u>Comb.</u> |
| 7. _____ $\text{Fe}_2\text{O}_3 +$ <u>3</u> $\text{H}_2 \rightarrow$ <u>2</u> $\text{Fe} +$ <u>3</u> H_2O | <u>S.R.</u> |
| 8. <u>4</u> $\text{FeS} +$ <u>3</u> $\text{O}_2 \rightarrow$ <u>2</u> $\text{Fe}_2\text{O}_3 +$ <u>4</u> SO_2 | <u>Comb.</u> |
| 9. <u>1</u> $\text{Cu}_2\text{S} +$ <u>3</u> $\text{O}_2 \rightarrow$ <u>2</u> $\text{Cu}_2\text{O} +$ <u>2</u> SO_2 | <u>Comb.</u> |
| 10. <u>4</u> $\text{NH}_3 +$ <u>5</u> $\text{O}_2 \rightarrow$ <u>4</u> $\text{NO} +$ <u>6</u> H_2O | <u>Comb.</u> |
| 11. _____ $\text{C}_3\text{H}_8 +$ <u>5</u> $\text{O}_2 \rightarrow$ <u>3</u> $\text{CO}_2 +$ <u>4</u> H_2O | <u>Comb</u> |
| 12. <u>2</u> $\text{AlPO}_4 +$ <u>3</u> $\text{Na}_2\text{S} \rightarrow$ _____ $\text{Al}_2\text{S}_3 +$ <u>2</u> Na_3PO_4 | <u>D.R.</u> |
| 13. _____ $\text{C}_{25}\text{H}_{32} +$ <u>38</u> $\text{O}_2 \rightarrow$ <u>26</u> $\text{H}_2\text{O} +$ <u>25</u> CO_2 | <u>Comb</u> |
| 14. <u>2</u> $\text{FeCl}_3 +$ <u>3</u> $\text{Na}_2\text{CO}_3 \rightarrow$ _____ $\text{Fe}_2(\text{CO}_3)_3 +$ <u>6</u> NaCl | <u>D.R.</u> |
| 15. _____ $\text{P}_2\text{O}_5 +$ <u>6</u> $\text{H}_2\text{O} \rightarrow$ <u>4</u> H_3PO_4 | <u>Syn.</u> |
| 16. <u>4</u> $\text{NH}_3 +$ <u>7</u> $\text{O}_2 \rightarrow$ <u>4</u> $\text{NO}_2 +$ <u>6</u> H_2O | <u>Comb.</u> |
| 17. <u>2</u> $\text{C}_7\text{H}_6\text{O}_2 +$ <u>15</u> $\text{O}_2 \rightarrow$ <u>14</u> $\text{CO}_2 +$ <u>6</u> H_2O | <u>Comb.</u> |

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Rajat Kalia, Manas Kalia



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