

25

NUCLEAR CHEMISTRY

Practice Problems

In your notebook, solve the following problems.

SECTION 25.1 NUCLEAR RADIATION

- What happens to the mass number and atomic number of an atom that undergoes beta decay?
- A nuclide of an element undergoes alpha particle decay. How do the atomic number and mass number of the particle change?
- Give the composition of the nucleus of the following isotopes.
 - $^{224}_{86}\text{Rn}$
 - $^{232}_{90}\text{Th}$
 - $^{209}_{83}\text{Bi}$
- Complete each of the following equations.
 - $^{238}_{92}\text{U} \rightarrow \text{?} + ^4_2\text{He}$
 - $^{235}_{92}\text{U} \rightarrow ^{231}_{90}\text{Th} + \text{?}$
 - $^{234}_{90}\text{Th} \rightarrow \text{?} + ^0_{-1}\text{e}$

SECTION 25.2 NUCLEAR TRANSFORMATIONS

- Write a nuclear equation for the following radioactive processes.
 - alpha decay of francium-223
 - electron capture by beryllium-7
 - beta emission by argon-37
 - positron emission by boron-12
- Complete the equations for these transmutation reactions.
 - $^{24}_{12}\text{Mg} + ^4_2\text{He} \rightarrow ^{28}_{14}\text{Si} + \text{?}$
 - $^{238}_{92}\text{U} + ^4_2\text{He} \rightarrow \text{?} + ^{238}_{90}\text{Th} + ^1_0\text{n}$
 - $^{240}_{94}\text{Pu} + ^4_2\text{He} \rightarrow \text{?} + ^4_2\text{He}$
 - $^{238}_{92}\text{U} \rightarrow ^{234}_{90}\text{Th} + \text{?} + ^4_2\text{He} + 4\text{?}$
 - $^4_2\text{He} + \text{?} \rightarrow ^{238}_{92}\text{U} + ^1_0\text{n} + 6^4_2\text{He} + 2\text{?}$
- Po-214 has a relatively short half-life of 164 s. How many seconds would it take for 8.0 g of this isotope to decay to 0.25 g?
- How many days does it take for 16 g of polonium-210 to decay to 1.0 g? The half-life of polonium-210 is 137 days.
- By approximately what factor would the mass of a sample of copper-66 decrease in 50 minutes? The half-life of copper-66 is 5.48 min.
- In 5.48 minutes, 1.00 g of argon-39 decays to leave only 0.12 g. What is the half-life of argon-39?

25 Nuclear Chemistry Answer Key

Daniela Niemeyer



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