

13-4 Practice Problems

1. What volume would be occupied by 100. g of oxygen gas at a pressure of 1.50 atm and a temperature of 25°C?

$$100. \text{ g } \text{O}_2 \times \frac{1 \text{ mol } \text{O}_2}{32 \text{ g } \text{O}_2} = 3.13 \text{ mol } \text{O}_2$$

$$PV = nRT$$

$$(1.50 \text{ atm}) V = (3.13 \text{ mol}) \left(0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \right) (298 \text{ K})$$

$$V = 51.1 \text{ L}$$

2. An air-filled balloon has a volume of 225 L at 0.94 atm and 25°C. Soon after, the pressure changes to 0.99 atm and the temperature changes to 0°C. What is the new volume of the balloon?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \frac{(0.94 \text{ atm})(225 \text{ L})}{298 \text{ K}} = \frac{(0.99 \text{ atm}) V_2}{273 \text{ K}}$$

$$V_2 = 2.0 \times 10^2 \text{ L}$$

3. A gas confined in a 515-cm³ container exerts a pressure of 107.4 kPa at 38.6°C. At what Celsius temperature will it exert a pressure of 635.7 kPa if it is placed into a

6. How many grams of argon would it take to fill a light bulb with a volume of 0.475 L at STP?

$$PV = nRT$$

$$(1) (0.475) = n (0.0821) (273)$$

$$n = 0.0212 \text{ mol Ar} \times \frac{40 \text{ g}}{1 \text{ mol}} = 0.848 \text{ g Ar}$$

7. Dry ice is carbon dioxide in the solid state. 1.28 grams of dry ice are placed into a 5.00 L evacuated chamber that is maintained at 35.1°C. What is the pressure in the chamber in kPa after all the dry ice has sublimed into CO₂ gas?

$$1.28 \text{ g CO}_2 \times \frac{1 \text{ mol}}{44 \text{ g}} = 0.0291 \text{ mol CO}_2$$

$$PV = nRT$$

$$P(5.00) = (0.0291)(0.0821)(308.1)$$

$$P = 0.147 \text{ atm} \times \frac{101.3 \text{ kPa}}{1 \text{ atm}} = 14.9 \text{ kPa}$$

8. A sample of Br₂ gas is loaded into an evacuated demonstration bottle at STP. The volume of the bottle is 0.25 L. How many moles of Br₂ gas will be contained

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Laurie Sorge



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Machine

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