

KEY

1. The combustion of 1.00 mol of sucrose, $C_{12}H_{22}O_{11}$, evolves 5.65×10^3 kJ of heat. A bomb calorimeter has a calibrated heat capacity of $1.23 \text{ kJ/}^\circ\text{C}$. How many grams of sucrose should be burned to raise the temperature of the calorimeter and its contents from 23.0°C to 80.0°C ? $\Delta T = 80.0^\circ - 23.0^\circ = 57.0^\circ\text{C}$

$$q = C_{cal} \cdot \Delta T = 1.23 \frac{\text{kJ}}{^\circ\text{C}} \cdot 57.0^\circ\text{C} = 70.11 \text{ kJ needed.}$$

$$1 \text{ mole sucrose} = 342.3 \text{ g} ; \frac{342.3 \text{ g}}{5650 \text{ kJ}} = \frac{x \text{ g}}{70.11 \text{ kJ}} ; x = \boxed{4.25 \text{ g sucrose}}$$

* don't have to do this.

The heat of reaction for burning 1 mole of a certain compound X is known to be -477.7 kJ . The calorimeter constant of the bomb being used is $2.5 \times 10^3 \text{ J/}^\circ\text{C}$ and the initial temperature of the water is 23.2°C .

- (A) If 22.54 g of compound X (MM = 46.0) is burned in the bomb calorimeter containing 2000 mL of water (S.H. = $4.184 \text{ J/g}^\circ\text{C}$), what will be the final temperature?

$$\boxed{T_f = 44.7^\circ\text{C}}$$

- (B) How much water can be warmed from 23.2°C to 56.5°C when 52.3 g of the compound is burned in the bomb?

$$\boxed{3300 \text{ g H}_2\text{O}}$$

3. A 500.0-g piece of metal at 60.0°C is placed in 200.0 g of water at 22.0°C contained in a coffee cup calorimeter. The metal and water come to the same temperature at 32.5°C . The specific heat of water is $4.184 \text{ J/g}^\circ\text{C}$. $q_{\text{lost by metal}} = q_{\text{gained by water}}$ $\Delta T_{\text{water}} = 32.5^\circ - 22.0^\circ = 10.5^\circ\text{C}$

- (A) How much heat energy did the metal give up to the water?

$$q_{\text{water}} = (200.0 \text{ g})(4.184 \text{ J/g}^\circ\text{C})(10.5^\circ\text{C}) = \boxed{8786 \text{ J released by metal}}$$

12 4 Review And Reinforcement Calorimetry Answers

Roger Porter



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