

The ABC's of Gas Laws

A

Avogadro's Law

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

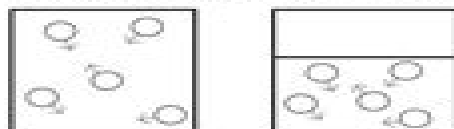
- Equal volume of gases, at same pressure and temperature, have an equal number of moles.
- Volume is directly proportional to the number of moles.
- Explanation of Gay-Lussac's Law of combining volumes.

B

Boyle's Law

$$P_1 V_1 = P_2 V_2$$

- The smaller the volume of the container, the greater the pressure on a gas due to more frequent collisions.
- The force of lots of collisions between gas molecules and walls of the container is responsible for the pressure exerted by a gas.
- Constant temperature and number of moles.



C

Charles's Law

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

- Constant pressure and number of moles.
- As the temperature of a gas increases, the gas's volume increases.
- As temperature increases, the speed of gas molecules increases, which makes the molecule's collisions with the walls of the container more energetic, thus increasing pressure. If the container is not rigid, then the volume of gas increases (Boyle's Law).

D

Dalton's Law of Partial Pressure

$$P_{\text{total}} = P_1 + P_2 + \dots P_n$$

$$P_i = X_i P_{\text{total}}$$

- Partial Pressure:** The pressure exerted by each gas in a mixture of ideal gases.
- The total pressure of a mixture of ideal gases is the sum of the partial pressures of all the gases in the mixture.

Total pressure of a gas mixture is determined by the number of moles of each gas present in the mixture, called it's **mole fraction**. The sum of the mole fractions for all the gases in a mixture should equal 1.

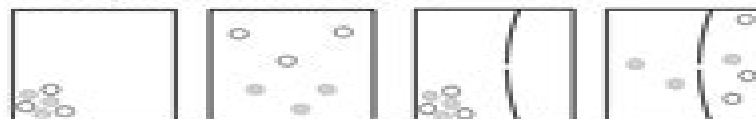
$$X_i = \frac{n_i}{n_{\text{total}}} \quad X_1 + X_2 = 1$$

E

Effusion (Graham's Law)

$$\frac{\text{Rate}_A}{\text{Rate}_B} = \frac{\text{Time}_B}{\text{Time}_A} = \sqrt{\frac{\text{Molar Mass}_B}{\text{Molar Mass}_A}}$$

- Effusion** - the relative rates at which a gas in a gas mixture escapes from a container through a hole.
- The larger the molecule, the smaller it's relative rate is. Larger molecules move slower so they cannot travel across a space as quickly (**diffusion**) or find the hole to escape as fast (**effusion**).



Diffusion

Effusion

G

Gay-Lussac's Law

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- As temperature increases, the pressure increases in a closed container due to more frequent collisions.
- Constant volume and number of moles.



12 3 Prentice Hall Gas Laws Section Review Answers

Mariam M. El-Awa



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Confidentiality in Arbitration Mariam M. El-Awa,2016-07-06 This book addresses the issue of privacy and confidentiality in the broader context of the Egyptian legal system The volume opens with an overview of the major approaches to confidentiality adopted in various jurisdictions It goes on to examine the duties of

confidentiality and privacy in arbitration law and practice on the basis of interviews with 30 law professors and practitioners who often act as arbitrators or counsel for parties in arbitral disputes together with the relevant Egyptian arbitration law provisions The book takes into account the relevant provisions in the arbitration laws of Syria Saudia Arabia and Yemen It moves on to explore the relation between arbitration and the judicial system and the extent to which the former should borrow its rules from the latter with regard to publicity and the rule of public trial Finally this book looks at the right to privacy as a a constitutional right as a potential basis for a legal duty of confidentiality in arbitration and the duties stemming from this constitutional right in the various laws of Egypt as well as b the constraints imposed on the right to privacy in particular those stemming from the constitutional principles of freedom of speech and freedom of the press The main conclusion is that confidentiality does indeed exist in arbitration However its legal basis is not the law on arbitration or the arbitration agreement It is in fact a corollary of the fundamental right to privacy granted in the Egyptian legal system to both natural and legal persons

Energy and the Wealth of Nations Charles A.S. Hall,Kent Klitgaard,2018-03-05 In this updated edition of a groundbreaking text concepts such as energy return on investment EROI provide powerful insights into the real balance sheets that drive our petroleum economy Hall and Klitgaard explore the relation between energy and the wealth explosion of the 20th century and the interaction of internal limits to growth found in the investment process and rising inequality with the biophysical limits posed by finite energy resources The authors focus attention on the failure of markets to recognize or efficiently allocate diminishing resources the economic consequences of peak oil the high cost and relatively low EROI of finding and exploiting new oil fields including the much ballyhooed shale plays and oil sands and whether alternative energy technologies such as wind and solar power can meet the minimum EROI requirements needed to run society as we know it For the past 150 years economics has been treated as a social science in which economies are modeled as a circular flow of income between producers and consumers In this perpetual motion of interactions between firms that produce and households that consume little or no accounting is given of the flow of energy and materials from the environment and back again In the standard economic model energy and matter are completely recycled in these transactions and economic activity is seemingly exempt from the Second Law of Thermodynamics As we enter the second half of the age of oil when energy supplies and the environmental impacts of energy production and consumption are likely to constrain economic growth this exemption should be considered illusory at best This book is an essential read for all scientists and economists who have recognized the urgent need for a more scientific empirical and unified approach to economics in an energy constrained world and serves as an ideal teaching text for the growing number of courses such as the authors own on the role of energy in society The Athenaeum ,1841 Operations Management Joel D.

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