

Atoms

1. What are the three fundamental particles that compose all atoms? Indicate the electrical charge and relative mass of each of these particles. Where is each type of particle found in the atom?

Name	Symbol	Charge	Mass (g)	Mass (amu)	Location
Proton	p^+	+1	1.673×10^{-24}	1	Nucleus
Neutron	n^0	0	1.675×10^{-24}	1	Nucleus
Electron	e^-	-1	9.109×10^{-28}	0.0005	Outside Nucleus

2. What are isotopes?

Usually all isotopes of a given element share the same number of protons, each isotope differs from the others in its number of neutrons.

3. What are ions? How are ions formed from atoms? To what do the terms cation and anion refer?

Very simply an ion is an atom or molecule that has gained or lost an electron. Both positive and negative ions exist.

4. Write the name and atomic number for each of the following elements:

- a. 15 **Phosphorus 15** b. 55 **Cesium 55** c. 15 **Phosphorus 15**
 d. 18 **Argon 18** e. 82 **Lead 82**

5. Write the name and chemical symbol for each of the following elements:

- a. 19 **K Potassium** c. 1 **H Hydrogen** e. 52 **Te Tellurium**
 b. 12 **Mg Magnesium** d. 6 **C Carbon** f. 2 **He Helium**

6. Indicate the number of protons, neutrons, and electrons in isolated atoms having the following nuclear symbols:

- a. ${}^4_2\text{He}$ **2 2 2** b. ${}^{37}_{17}\text{Cl}$ **17 20 17** c. ${}^{60}_{20}\text{Ca}$ **20 40 20**

7. Which of the subatomic particles is responsible for the chemical behavior of a given type of atom? Why?

Electrons An electron generates an electric field that exerts an attractive force on a particle with a positive charge, such as the proton, and a repulsive force on a particle with a negative charge.

8. Write the name and atomic number for each of the following elements:

- a. 31 **P Phosphorus** c. 7 **N Nitrogen** e. 16 **S Sulfur**
 b. 6 **C Carbon** d. 35 **Br Bromine** f. 24 **Cr Chromium**

9. Write the name and chemical symbol for each of the following elements:

- a. 36 **Krypton** c. 15 **P Phosphorus** e. 29 **Cu Copper**
 b. 92 **U Uranium** d. 79 **Au Gold** f. 8 **O Oxygen**

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Camilla Rothe

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name Camilla Rothe.

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if x is not a whole number it is best to leave the answer as a fraction or in mixed numeral form rather than as a decimal because the decimal answer would often have to be rounded x 5 16 24 x 5 16 24 24 16 16 24 384 x 5 384 16 24 multiply both sides by the product of the denomiinators

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