

Chapter 11 – Thermochemistry – Heat and Chemical Change

Chapter 11: 1 – 35, 52, 60, 61, 71

Section 11.1 – The Flow of Energy - Heat

Practice Problems

1. When 435 J of heat is added to 3.4 g of olive oil at 21°C, the temperature increases to 85°C. What is the specific heat of olive oil?

Knowns: $q = 435 \text{ J}$; $m_{\text{olive oil}} = 3.4 \text{ g}$; $\Delta T = (85^\circ\text{C} - 21^\circ\text{C})$

Use $C = q / m \cdot \Delta T$

$$C = \frac{435 \text{ J}}{3.4 \text{ g} \cdot 64^\circ\text{C}} = 2.0 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$$

2. A 1.55-g piece of stainless steel absorbs 141 J of heat when its temperature increases by 178°C. What is the specific heat of the stainless steel?

Knowns: $m_{\text{stainless steel}} = 1.55 \text{ g}$; $q = 141 \text{ J}$; $\Delta T = 178^\circ\text{C}$

Use $C = q / m \cdot \Delta T$

$$C = \frac{141 \text{ J}}{1.55 \text{ g} \cdot 178^\circ\text{C}} = 0.511 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$$

3. How much heat is required to raise the temperature of 250.0 g of mercury 52°C?

Knowns: $m_{\text{Hg}} = 250.0 \text{ g}$; $\Delta T = 52^\circ\text{C}$; $C_{\text{Hg}} = 0.14 \text{ J/g} \cdot ^\circ\text{C}$

Use $C = q / m \cdot \Delta T$; rearrange to solve for q : $q = C \cdot m \cdot \Delta T$

$$q = \left(0.14 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}\right) (250.0 \text{ g})(52^\circ\text{C}) = 1,820 \text{ J}$$

Section Review 11.1

4. Define energy and explain how energy and heat are related.

Energy is the capacity for doing work or supplying heat. Heat is energy that is transferred between objects, when their temperatures are different.

5. Explain the difference between heat capacity and specific heat.

The specific heat of a substance is independent of its mass, whereas the heat capacity of an object is proportional to its mass.

6. Will the specific heat of 50 g of a substance be the same as, or greater than, the specific heat of 10 g of the same substance? Explain.

The specific heat between the two substances will be the same. Specific heat is not dependent upon the mass, whereas heat capacity is.

7. On a sunny day, why does the concrete deck around an outdoor swimming pool become hot, while the water stays cool?

The water stays cooler because it has a higher heat capacity than concrete. The sun's heat raised the temperature of the concrete more than that of the water.

11 Thermochemistry Heat And Chemical Change Answers

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Table of Contents **11 Thermochemistry Heat And Chemical Change Answers**

1. Understanding the eBook **11 Thermochemistry Heat And Chemical Change Answers**
 - The Rise of Digital Reading **11 Thermochemistry Heat And Chemical Change Answers**
 - Advantages of eBooks Over Traditional Books
2. Identifying **11 Thermochemistry Heat And Chemical Change Answers**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **11 Thermochemistry Heat And Chemical Change Answers**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **11 Thermochemistry Heat And Chemical Change Answers**
 - Personalized Recommendations

- 11 Thermochemistry Heat And Chemical Change Answers User Reviews and Ratings
- 11 Thermochemistry Heat And Chemical Change Answers and Bestseller Lists
- 5. Accessing 11 Thermochemistry Heat And Chemical Change Answers Free and Paid eBooks
 - 11 Thermochemistry Heat And Chemical Change Answers Public Domain eBooks
 - 11 Thermochemistry Heat And Chemical Change Answers eBook Subscription Services
 - 11 Thermochemistry Heat And Chemical Change Answers Budget-Friendly Options
- 6. Navigating 11 Thermochemistry Heat And Chemical Change Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 11 Thermochemistry Heat And Chemical Change Answers Compatibility with Devices
 - 11 Thermochemistry Heat And Chemical Change Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 11 Thermochemistry Heat And Chemical Change Answers
 - Highlighting and Note-Taking 11 Thermochemistry Heat And Chemical Change Answers
 - Interactive Elements 11 Thermochemistry Heat And Chemical Change Answers
- 8. Staying Engaged with 11 Thermochemistry Heat And Chemical Change Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 11 Thermochemistry Heat And Chemical Change Answers
- 9. Balancing eBooks and Physical Books 11 Thermochemistry Heat And Chemical Change Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 11 Thermochemistry Heat And Chemical Change Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 11 Thermochemistry Heat And Chemical Change Answers
 - Setting Reading Goals 11 Thermochemistry Heat And Chemical Change Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 11 Thermochemistry Heat And Chemical Change Answers
 - Fact-Checking eBook Content of 11 Thermochemistry Heat And Chemical Change Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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