## Multiple Choice

Identify the choice that best completes the statement or answers the question.

- The only sure evidence for a chemical reaction is a. the formation of a gas.

  - b. changes in properties.
  - the production of one or more new substances.
- When a chemical reaction takes place in an open system.
  - matter cannot move at all.
  - b. matter can enter from or escape to the surroundings.
  - matter is not allowed to enter from or escape to the surroundings.
- 3. Which of the following is a balanced chemical equation?
  - a. 2 Fe<sub>2</sub>O<sub>1</sub> + 3 C → 4 Fe + 3 CO<sub>2</sub>
  - b. H<sub>2</sub>O<sub>2</sub> → H<sub>2</sub>O + O<sub>2</sub>
  - c. SO<sub>2</sub> + O<sub>2</sub> + 2 H<sub>2</sub>O → 4 H<sub>2</sub>SO<sub>4</sub>
- 4. What type of reaction is FeS + 2 HCl → FeCl<sub>2</sub> + H<sub>2</sub>S?
  - a single replacement reaction
  - b. a double replacement reaction
  - a synthesis reaction
- 5. What is the best form of fire safety?
  - a. fire prevention
  - b. the fire department
  - c. an extinguisher
- Anything that has mass and takes up space is considered
  - a. energy.
  - b. chemistry.
  - c. matter.
- 7. The best way to put out a small kitchen fire is to
  - a. blow it out.
  - b. turn off the appliances.
  - c cover it or use a home fire extinguisher.
- 8. A chemical reaction that absorbs energy in the form of heat is described as
  - a. combustion.
  - b. unbalanced.
  - endothermie.

# 11 Chemical Reactions Practice Problems Answers 11 1

**Sidney William Benson** 

#### 11 Chemical Reactions Practice Problems Answers 11 1:

General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie s classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works. This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value **Calculating Theoretical Yield Is Not** Impossible! Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as practice problems and questions to ensure mastery Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory Addresses the necessity of balancing the reaction equation the centrality of the mole in these calculations and the role of significant figures in reporting the answer Explains how to approach the calculations when using neat liquids or solutions The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product All of the problems in the book include detailed solutions with accompanying text to explain the answers and ancillaries also include suggestions for further reading The Latest and Best of TESS ,1991

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