

Chemistry Multiple Choice: Reactions

1. What is the general form of single displacement reactions?

- A) $A + B \rightarrow AB$ C) $AB + C \rightarrow CB + A$
B) $C_xH_y + O_2 \rightarrow CO_2 + H_2O$ D) $AB + CD \rightarrow AD + CB$

ANS: C

1. In a single displacement reaction, aqueous iron(III) sulphate reacts with barium powder. What is the balanced chemical equation for this reaction?

- A) $Fe_2(SO_4)_3(aq) + Ba(s) \rightarrow BaSO_4(s) + Fe(s)$
B) $Fe_2(SO_4)_3(aq) + 3Ba(s) \rightarrow 3BaSO_4(s) + 2Fe(s)$
C) $Fe_2(SO_4)_3(aq) + 2Ba(s) \rightarrow 2BaSO_4(s) + 3Fe(s)$
D) $FeSO_4(aq) + Ba(s) \rightarrow BaSO_4(s) + Fe(s)$

ANS: B

1. What would be the general chemical equation for decomposition reactions?

- A) $C \rightarrow A + B$ C) $AB \rightarrow A + B$
B) $AB + C \rightarrow ABC$ D) $A + O \rightarrow AOH$

ANS: C

1. What type of reaction almost always gives off large amounts of heat and light?

- A) combustion B) neutralization C) decomposition D) synthesis

ANS: A

1. What are the reactants of every hydrocarbon combustion reaction?

- A) $O_2 + H_2O$ B) $C_xH_y + O_2$ C) $C_xH_yO_x + O_2$ D) $CO_2 + H_2O$

ANS: B

1. The following balanced equation represents the reaction of aluminum with copper(II) chloride.



This is an example of a:

- A) formation reaction B) combustion reaction
C) single displacement reaction D) double displacement reaction

ANS: C

1. The products of the reaction of aqueous sodium iodide, $NaI_{(aq)}$, and aqueous lead(II) nitrate, $Pb(NO_3)_{2(aq)}$, are:

- A) $NaPb_{(s)}$ and $I_2NO_{3(aq)}$ B) $Na_{(s)}$ and $I_2NO_{3(aq)}$
C) $NaNO_{3(aq)}$ and $PbI_{2(s)}$ D) $NaNO_3(aq)$ and $PbI_{2(s)}$

ANS: D

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