

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

- Classify a solution as neutral, acidic, or basic, given the hydrogen-ion or hydroxide-ion concentration
- Convert hydrogen-ion concentrations into values of pH and hydroxide-ion concentrations into values of pOH
- Describe the purpose of pH indicators

Vocabulary

- self-ionization
- ion-product constant for water (K_w)
- acidic solution
- basic solution
- alkaline solutions
- pH

Key Equations

- $K_w = [\text{H}^+] \times [\text{OH}^-] = 1.0 \times 10^{-14} \text{M}^2$
- $\text{pH} = -\log [\text{H}^+]$
- $\text{pOH} = -\log [\text{OH}^-]$
- $\text{pH} + \text{pOH} = 14$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Water molecules can _____ 1 _____ to form hydrogen ions (H^+) and hydroxide ions (OH^-). The concentrations of these ions in pure water at 25°C are both equal to _____ 2 _____ mol/L.

The pH scale, which has a range from _____ 3 _____, is used to denote the _____ 4 _____ concentration of a solution. On this scale, 0 is strongly _____ 5 _____, 14 is strongly _____ 6 _____, and 7 is _____ 7 _____. Pure water at 25°C has a pH of _____ 8 _____.

The _____ 9 _____ constant for water has a value of 1.0×10^{-14} . Thus, the product of the concentrations of _____ 10 _____ ions and _____ 11 _____ ions in aqueous solution will always equal 1.0×10^{-14} .

1. ionize
2. 1×10^{-7} mol/L
3. 0-14
4. H^+
5. acidic
6. basic
7. neutral
8. 7
9. ion product
10. $\text{H}^+ / \text{H}_3\text{O}^+$
11. OH^-

19 3 Section Review Answers

Harold Torrance



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