

# ANSWER KEY: Ch. 17 Acid-base and Solubility Equilibria

**Question 1** Which of the following acids is not a weak acid? *1. Which of the following acids is not a weak acid?*

**Answer 1** *1. Which of the following acids is not a weak acid?*

The strong, nonvolatile acids and weak acids are listed in concordance with acid-base literature. Which one of the following statements is not true?

- ☐ A. The equivalence point and end point will always be exactly the same for any titration.
- ☐ B. As the concentration of the analyte or titrant in the solution being titrated is the same as the number of equivalents of the titrant added.
- ☐ C. The end point of a titration is often dependent on an indicator color change.
- ☐ D. The equivalence point of a titration can be found using a pH meter.

**Question 2**

*2. How many milliliters of 0.10 M NaOH are needed to neutralize 10 grams of HCl dissolved in water? (HCl is a strong acid.)*

**Answer 2** *2. How many milliliters of 0.10 M NaOH are needed to neutralize 10 grams of HCl dissolved in water? (HCl is a strong acid.)*

- ☐ A. 1000 mL
- ☐ B. 100 mL
- ☐ C. 250 mL
- ☐ D. 25 mL

**Question 3** *3. Which of the following is a weak acid?*

**Answer 3** *3. Which of the following is a weak acid?*

- ☐ A.  $1.0 \times 10^{-2} \text{ mol/L}$
- ☐ B.  $1.0 \times 10^{-2} \text{ mol/L}$

# 17 1 Answer Key Solubility Equilibria

**Disha Experts**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

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**Introductory Qualitative Analysis** Jacob Cornog, Warren Chase Vosburgh, 1928

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