



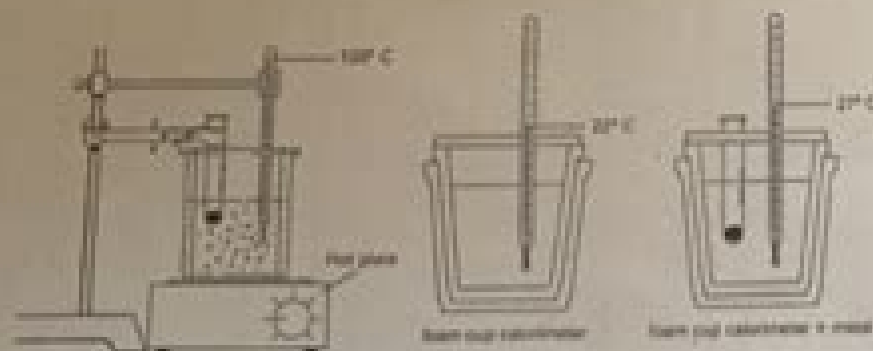
INTERPRETING GRAPHICS

USE WITH SECTION 11.1

A student performed an experiment to determine the specific heat of an unknown metal. The data she collected is reproduced in the table below. Use this information to answer the following questions.

| Quantity | Trial 1 | Trial 2 |
|---------------------------------|----------|----------|
| 1. Mass of test tube + metal | 118.19 g | 118.21 g |
| 2. Mass of test tube | 47.88 g | 47.88 g |
| 3. Mass of metal | 70.31 g | 70.33 g |
| 4. Mass of calorimeter | 7.237 g | 7.218 g |
| 5. Mass of calorimeter + water | 46.127 g | 45.779 g |
| 6. Mass of water | 38.890 g | 38.562 g |
| 7. Initial temperature of metal | 100.0 °C | 100.0 °C |
| 8. Initial temperature of water | 22.8 °C | 21.6 °C |
| 9. Final temperature of water | 27.8 °C | 26.3 °C |

Please
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all
work!!



1. What was the final temperature of the metal?

a. Trial 1

b. Trial 2

2. What was the ΔT for the metal?

a. Trial 1

b. Trial 2

3. What was the ΔT for the water?

a. Trial 1

b. Trial 2

4. Calculate the heat change for the water. (Specific heat capacity for water is $4.184 \text{ J/g}^\circ\text{C}$)

a. Trial 1

b. Trial 2

5. Calculate the heat change for the metal.

a. Trial 1

b. Trial 2

6. Calculate the specific heat of the metal.

a. Trial 1

b. Trial 2

7. What metal might this have been?

a. aluminum

b. silver

c. iron

d. mercury

88 Review Module 1 (Chapters 8-12)

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AW Rasmussen



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