12 • The Gas Laws

THE COMBINED GAS LAW

In practical terms, it is often difficult to hold any of the variables constant. When there is a change in pressure, volume and temperature, the combined gas law is used.

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$
 or $P_1 \times V_1 \times T_2 = P_2 \times V_2 \times T_1$
 $K = {}^{\circ}C + 273$

Ser atet

Complete the following chart.

	P ₁	$\mathbf{V_1}$	T ₁	P ₂	V_2	T ₂
1	1.50 atm	3.00 L	293 K 20.0°C	2.50 atm		3051€
2	720. torr	256. mL	298 K		250. mL	3 23 K -50.0 °C
3	600. mmHg	2.50 L	325 K -22.0°€	760. mmHg	1.80 L	
4		750. mL	273 ℃ -0.00°C	2.00 atm	500. mL	248 K 25.0°C
5	95.0 kPa	4.00 L		101. kPa	6.00 L	471. K or 498. °C
6	650. torr		373 € -100.°C	900. torr	225. mL	423 K 150.°€
7	850. mmHg	1.50 L	288 K -15.0°€		2.50 L	303 K
8	125. kPa	125. mL		100. kPa	100 mL	348 K

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